

Series : RPQ3S



SET~1

प्रश्न-पत्र कोड
Q.P. Code

56/3/1

रोल नं.

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Roll No.

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परीक्षार्थी प्रश्न-पत्र कोड को उत्तर-पुस्तिका के मुख-पृष्ठ पर अवश्य लिखें।
Candidates must write the Q.P. Code on the title page of the answer-book.

नोट	NOTE
(I) कृपया जाँच कर लें कि इस प्रश्न-पत्र में मुद्रित पृष्ठ 19 हैं।	(I) Please check that this question paper contains 19 printed pages.
(II) प्रश्न-पत्र में दाहिने हाथ की ओर दिए गए प्रश्न-पत्र कोड को परीक्षार्थी उत्तर-पुस्तिका के मुख-पृष्ठ पर लिखें।	(II) Q.P. Code given on the right hand side of the question paper should be written on the title page of the answer-book by the candidate.
(III) कृपया जाँच कर लें कि इस प्रश्न-पत्र में 33 प्रश्न हैं।	(III) Please check that this question paper contains 33 questions.
(IV) कृपया प्रश्न का उत्तर लिखना शुरू करने से पहले, उत्तर-पुस्तिका में यथा स्थान पर प्रश्न का क्रमांक अवश्य लिखें। []	(IV) Please write down the Serial Number of the question in the answer-book at the given place before attempting it.
(V) इस प्रश्न-पत्र को पढ़ने के लिए 15 मिनट का समय दिया गया है। प्रश्न-पत्र का वितरण पूर्वाह्न में 10.15 बजे किया जाएगा। 10.15 बजे से 10.30 बजे तक परीक्षार्थी केवल प्रश्न-पत्र को पढ़ेंगे और इस अवधि के दौरान वे उत्तर-पुस्तिका पर कोई उत्तर नहीं लिखेंगे।	(V) 15 minute time has been allotted to read this question paper. The question paper will be distributed at 10.15 a.m. From 10.15 a.m. to 10.30 a.m., the candidates will read the question paper only and will not write any answer on the answer-book during this period.



रसायन विज्ञान (सैद्धान्तिक)
CHEMISTRY (Theory)



निर्धारित समय : 3 घण्टे

Time allowed : 3 hours

अधिकतम अंक : 70

Maximum Marks : 70



General Instructions :

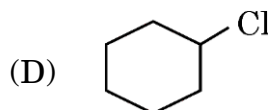
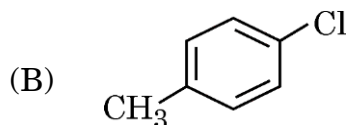
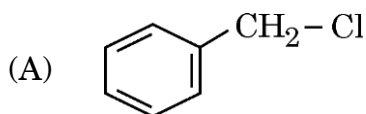
Read the following instructions carefully and follow them :

- (i) This question paper contains **33** questions. **All** questions are **compulsory**.
- (ii) This question paper is divided into **five** sections – **Section A, B, C, D and E**.
- (iii) **Section A** – questions number **1 to 16** are multiple choice type questions. Each question carries **1** mark.
- (iv) **Section B** – questions number **17 to 21** are very short answer type questions. Each question carries **2** marks.
- (v) **Section C** – questions number **22 to 28** are short answer type questions. Each question carries **3** marks.
- (vi) **Section D** – questions number **29 and 30** are case-based questions. Each question carries **4** marks.
- (vii) **Section E** – questions number **31 to 33** are long answer type questions. Each question carries **5** marks.
- (viii) There is no overall choice given in the question paper. However, an internal choice has been provided in few questions in all the sections except Section A.
- (ix) Kindly note that there is a separate question paper for Visually Impaired candidates.
- (x) Use of calculator is **not** allowed.

SECTION A

Questions no. **1 to 16** are Multiple Choice type Questions, carrying **1** mark each. $16 \times 1 = 16$

1. Which of the following will be the least reactive towards nucleophilic substitution reaction ?



2. The intermediate formed during the slowest step involved in the dehydration of alcohol is :

(A) protonated alcohol

(B) carbanion

(C) free radical

(D) carbocation

3. The oxidation number of Co in $[\text{Co}(\text{en})_3]_2(\text{SO}_4)_3$ is :

(A) + 3

(B) + 2

(C) + 4

(D) + 6





4. According to Werner's theory, the primary valencies of the central metal atom :
- are satisfied by neutral molecules or negative ions.
 - are equal to its coordination number.
 - are satisfied by negative ions.
 - are non-ionisable.
5. Which of the following reagent is used to distinguish between $(C_2H_5)_2NH$ and $(C_2H_5)_3N$?
- $CHCl_3 + KOH$
 - $C_6H_5SO_2Cl$
 - Conc. $HCl + ZnCl_2$
 - $NaOH + I_2$
6. Among the following, which is the strongest base ?
- -
 -
 -
7. In aqueous solution, $Cr_2O_7^{2-}$ ion converts to which of the following in alkaline medium ?
- Cr^{3+}
 - CrO_4^{2-}
 - CrO
 - CrO_3
8. The boiling point of an azeotropic mixture of water and ethanol is less than that of pure water and ethanol. The mixture shows :
- positive deviation from Raoult's Law.
 - negative deviation from Raoult's Law.
 - no deviation from Raoult's Law.
 - that the solution is an ideal solution.





9. The mole fraction of a solute in 2.0 molal aqueous solution is :
 (A) 1.87 (B) 0.347
 (C) 0.0347 (D) 0.00347
10. On electrolysis of very dilute aqueous solution of NaCl using platinum electrodes :
 (A) H₂ gas is evolved at anode.
 (B) Na is produced at cathode.
 (C) O₂ gas is evolved at anode.
 (D) H₂ gas is evolved at cathode.
11. For a reaction $3A \rightarrow 2B$, the rate of reaction $\frac{+d[B]}{dt}$ is equal to :
 (A) $-\frac{3}{2} \frac{d[A]}{dt}$ (B) $-\frac{2}{3} \frac{d[A]}{dt}$
 (C) $-\frac{1}{3} \frac{d[A]}{dt}$ (D) $+2 \frac{d[A]}{dt}$
12. On hydrolysis, which of the following carbohydrates gives only glucose ?
 (A) Sucrose (B) Galactose
 (C) Lactose (D) Maltose

For Questions number 13 to 16, two statements are given — one labelled as Assertion (A) and the other labelled as Reason (R). Select the correct answer to these questions from the codes (A), (B), (C) and (D) as given below :

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
 (B) Both Assertion (A) and Reason (R) are true, but Reason (R) is **not** the correct explanation of the Assertion (A).
 (C) Assertion (A) is true, but Reason (R) is false.
 (D) Assertion (A) is false, but Reason (R) is true.
13. **Assertion (A) :** Separation of Zr and Hf is difficult.

Reason (R) : Because Zr and Hf lie in the same group of the periodic table.





- 14.** *Assertion (A) :* Phenol is less acidic than 4-methylphenol.
Reason (R) : The presence of an electron releasing group in phenol makes it less acidic.
- 15.** *Assertion (A) :* Aromatic primary amines cannot be prepared by Gabriel phthalimide synthesis.
Reason (R) : Gabriel phthalimide synthesis is used for the preparation of primary aliphatic amines.
- 16.** *Assertion (A) :* Order of reaction is applicable to elementary as well as complex reactions.
Reason (R) : Order of a reaction is an experimental quantity.

SECTION B

- 17.** For the compound having molecular formula C_4H_9Br , write :
 (a) the isomer which is most reactive towards S_N1 displacement.
 (b) the isomer which, on reacting with Na metal in the presence of dry ether, gives 2,5-Dimethylhexane. $1+1=2$
- 18.** For a reaction $A + B \rightarrow \text{Products}$, the rate law is :

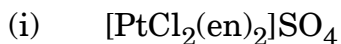
$$\text{Rate} = k[A][B]^{\frac{3}{2}}$$
 Write the overall order of the reaction. Can this reaction be an elementary reaction ? Give reason in support of your answer. 2
- 19.** Explain why, on addition of 1 mol of KCl to 1 litre of water, the boiling point of water increases, while the addition of 1 mol of methyl alcohol to 1 litre of water decreases the boiling point. 2





20. (a) Write IUPAC names of the following compounds :

1+1=2



OR

(b) Define the following terms with a suitable example in each case : 1+1=2

(i) Ambidentate ligand

(ii) Double salt

21. How do you explain the following ?

1+1=2

(a) Presence of a straight chain in glucose

(b) Presence of five – OH groups in glucose which are attached to different carbon atoms

SECTION C

22. The half-life period of a radioactive element is 1.5×10^{10} years. Calculate the time in which the activity of the element is reduced to 75% of its original value.

3

[Given : $\log 2 = 0.30$, $\log 3 = 0.48$, $\log 4 = 0.60$]

23. State Kohlrausch's law of independent migration of ions. With the help of a curve, explain why it is not easy to determine Λ_m° for weak electrolytes by extrapolating the concentration – molar conductivity curve, as it is for strong electrolytes.

3

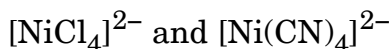
24. An antifreeze solution is prepared by dissolving 31 g of ethylene glycol (Molar mass = 62 g mol^{-1}) in 600 g of water. Calculate the freezing point of the solution. (K_f for water = $1.86 \text{ K kg mol}^{-1}$)

3





25. (a) Answer the following questions about the complexes 3×1=3



- (i) Write the hybridization involved in each case.
- (ii) Which of them is the inner orbital complex and which one is the outer orbital complex ?
- (iii) Compare their magnetic behaviour.

[Atomic number : Ni = 28]

OR

- (b) (i) Name two coordination compounds which are important in biological systems.
- (ii) What is meant by chelate effect ? Give an example.
- (iii) Why are low spin tetrahedral complexes rarely formed ? 3×1=3

26. Give reasons for the following : 3×1=3

- (a) Although chlorine is an electron-withdrawing group, yet it is ortho-, para-directing in electrophilic aromatic substitution reactions.
- (b) Aryl halides cannot be prepared by reacting phenol with concentrated halogen acids or phosphorus halides.
- (c) Chloroform is stored in closed dark coloured bottles.

27. How will you obtain the following from benzenediazonium chloride ? Give chemical equations involved : 3×1=3

- (a) Chlorobenzene
- (b) Benzene
- (c) Benzonitrile

28. An organic compound (A) with molecular formula $\text{C}_3\text{H}_5\text{N}$ on reaction with $\text{C}_6\text{H}_5\text{MgBr}$ followed by hydrolysis, gives a compound (B). Compound (B) forms an orange-red precipitate with 2,4-DNP reagent and does not give iodoform test. It neither reduces Tollens' or Fehling's reagent nor does it decolourise bromine water. On drastic oxidation with chromic acid it gives a carboxylic acid (C) having molecular formula $\text{C}_7\text{H}_6\text{O}_2$. Identify the compounds (A), (B) and (C). Write the reactions of compound (A) with $\text{C}_6\text{H}_5\text{MgBr}$ followed by hydrolysis to give compound (B). 3



**SECTION D**

The following questions are case-based questions. Read the case carefully and answer the questions that follow.

29. Living systems are made up of various complex biomolecules like carbohydrates, proteins, nucleic acids, lipids, etc. Proteins and carbohydrates are essential constituents of our food. In addition, some simple molecules like vitamins and mineral salts also play an important role in the functions of organisms. All proteins are polymers of α -amino acids. Proteins can be classified into two types on the basis of their molecular shape — Fibrous and Globular proteins. Vitamins are accessory food factors required in the diet. They are classified as fat-soluble and water-soluble. Deficiency of vitamins leads to many diseases. Nucleic acids are the polymers of nucleotides which in turn consist of a base, a pentose sugar and phosphate moiety. There are two types of nucleic acids — DNA and RNA. Nucleic acids are responsible for the transfer of characters from parents to offsprings.

- (a) Write the name of basic building units of proteins and nucleic acids. How can you differentiate between Fibrous and Globular proteins on the basis of their structures ? 2
- (b) (i) What products would be formed when a nucleotide from DNA containing thymine is hydrolyzed ? 1

OR

- (b) (ii) Write one structural difference between DNA and RNA. 1
- (c) Give one example each of a fat-soluble vitamin and a water-soluble vitamin. 1

30. Alcohols have higher boiling points than other compounds, namely hydrocarbons, ethers and haloalkanes of comparable molecular masses. On oxidation, primary alcohols yield aldehydes with mild oxidizing agents and carboxylic acids with strong oxidizing agents. Secondary alcohols yield ketones on oxidation while tertiary alcohols are resistant to oxidation. Ethers may be prepared by dehydration of alcohols and Williamson synthesis. The C – O bond in ethers can be cleaved by hydrogen halides. The presence of –OH group in phenols activates the aromatic ring towards electrophilic substitution and directs the incoming group to ortho and para positions due to resonance effect. In presence of NaOH, phenol generates phenoxide ion which is even more reactive than phenol. Thus, in alkaline medium, phenol undergoes Kolbe's reaction.

- (a) Name the reagents used in the following reactions : 2
- (i) Oxidation of a primary alcohol to aldehyde
- (ii) Oxidation of a primary alcohol to carboxylic acid

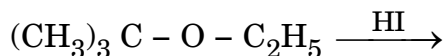




- (b) Write the reaction involved in Kolbe's reaction. 1
- (c) (i) Why are tertiary alcohols resistant to oxidation ? 1

OR

- (c) (ii) Write the products of the following reaction : 1



SECTION E

31. (a) (i) Describe giving reason which one of the following pairs has the property indicated :
- (I) Fe or Cu – higher melting point
- (II) Ti^{3+} or Sc^{3+} – coloured in aqueous solution
- (III) Cr or Zn – higher third ionisation enthalpy
- (ii) Write the ionic equations for the oxidizing action of MnO_4^- in acidic medium with :
- (I) Fe^{2+} ion
- (II) I^- ion 3+2=5

OR

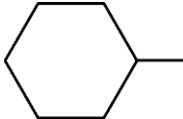
- (b) (i) A black-brown coloured solid (A) when fused with KOH in the presence of air, produces a dark green coloured compound (B) which on electrolytic oxidation in alkaline medium gives a dark purple coloured compound (C). Identify (A), (B) and (C). Write the reactions involved.
- (ii) What happens when an acidic solution of the green compound (B) is allowed to stand for some time ? Also write the equation involved. What is this type of reaction called ? 3+2=5

32. (a) (i) Write the product(s) when :
- (I) One mol of ethanal is treated with 1 mol of CH_3OH in the presence of dry HCl gas.
- (II) Benzaldehyde is treated with conc. NaOH.
- (III) Ethanoic acid is heated in the presence of P_2O_5 .
- (ii) Write a simple chemical test to distinguish between Ethanal and Propanal.
- (iii) Write the name of the reagent to transform Allyl alcohol to Propenal. 3+1+1=5

OR





- (b) (i) Draw the structure of the semicarbazone of acetone.
- (ii) Why are α -hydrogen atoms of aldehydes and ketones acidic in nature ?
- (iii) Arrange the following compounds in increasing order of their reactivity towards HCN :
 CH_3COCH_3 , CH_3CHO , $(\text{CH}_3)_3\text{C} - \underset{\text{O}}{\underset{\parallel}{\text{C}}} - \text{CH}_3$
- (iv) Write the reaction involved in Etard reaction.
- (v) Write the product when  reacts with Zn(Hg)/conc. HCl . 5×1=5

33. (a) (i) Calculate the electrode potential of a half-cell for zinc electrode dipping in 0.01 M ZnSO_4 solution at 25°C .
 Given : $E^\circ_{\text{Zn}^{2+}/\text{Zn}} = -0.76 \text{ V}$
 $[\log 10 = 1]$
- (ii) Write anode, cathode and overall reaction involved in dry cell.
- (iii) Equilibrium constant (K_c) is related to E°_{cell} , but not to E_{cell} .
 Why ? 2+2+1=5

OR

- (b) (i) The conductivity of 0.001 M solution of acetic acid is $3.905 \times 10^{-5} \text{ S cm}^{-1}$. Calculate its molar conductivity and degree of dissociation (α).
 Given : $\lambda^\circ_{\text{CH}_3\text{COO}^-} = 40.9 \text{ S cm}^2 \text{ mol}^{-1}$
 $\lambda^\circ_{\text{H}^+} = 349.6 \text{ S cm}^2 \text{ mol}^{-1}$
- (ii) Give reasons for the following :
- (I) Why does a mercury cell deliver a constant voltage for its entire life ?
- (II) Why is it necessary to use salt bridge in a galvanic cell ?

