

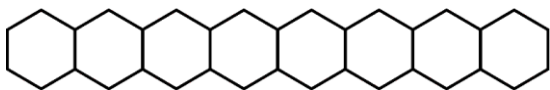


Series : WYXZ6

SET ~ 2

रोल नं.

Roll No.



प्रश्न-पत्र कोड

Q.P. Code

56/6/2

परीक्षार्थी प्रश्न-पत्र कोड को उत्तर-पुस्तिका के मुख-पृष्ठ पर अवश्य लिखें।

Candidates must write the Q.P. Code on the title page of the answer-book.

(I) कृपया जाँच कर लें कि इस प्रश्न-पत्र में मुद्रित पृष्ठ **23** हैं।

Please check that this question paper contains **23** printed pages.



(II) प्रश्न-पत्र में दाहिने हाथ की ओर दिए गए प्रश्न-पत्र कोड को परीक्षार्थी उत्तर-पुस्तिका के मुख-पृष्ठ पर लिखें।

Q.P. Code given on the right hand side of the question paper should be written on the title page of the answer-book by the candidate.

(III) कृपया जाँच कर लें कि इस प्रश्न-पत्र में **33** प्रश्न हैं।

Please check that this question paper contains **33** questions.

(IV) कृपया प्रश्न का उत्तर लिखना शुरू करने से पहले, उत्तर-पुस्तिका में यथा स्थान पर प्रश्न का क्रमांक अवश्य लिखें।

Please write down the Serial Number of the question in the answer-book at the given place before attempting it.

(V) इस प्रश्न-पत्र को पढ़ने के लिए 15 मिनट का समय दिया गया है। प्रश्न-पत्र का वितरण पूर्वाह्न में 10.15 बजे किया जाएगा। 10.15 बजे से 10.30 बजे तक परीक्षार्थी केवल प्रश्न-पत्र को पढ़ेंगे और इस अवधि के दौरान वे उत्तर-पुस्तिका पर कोई उत्तर नहीं लिखेंगे।

15 minute time has been allotted to read this question paper. The question paper will be distributed at 10.15 a.m. From 10.15 a.m. to 10.30 a.m., the candidates will read the question paper only and will not

write any answer on the answer-book during this period.



रसायन विज्ञान (सैद्धान्तिक)

CHEMISTRY (Theory)



निर्धारित समय : 3 घण्टे

Time allowed : 3 hours

अधिकतम अंक : 70

Maximum Marks : 70

56/6/2

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P.T.O.



General Instructions :

Read the following instructions carefully and follow them :

- (i) This question paper contains **33** questions. **All** questions are **compulsory**.
- (ii) This question paper is divided into **five** sections – **Section A, B, C, D and E**.
- (iii) **Section A** – questions number **1 to 16** are multiple choice type questions. Each question carries **1** mark.
- (iv) **Section B** – questions number **17 to 21** are very short answer type questions. Each question carries **2** marks.
- (v) **Section C** – questions number **22 to 28** are short answer type questions. Each question carries **3** marks.
- (vi) **Section D** – questions number **29 and 30** are case-based questions. Each question carries **4** marks.
- (vii) **Section E** – questions number **31 to 33** are long answer type questions. Each question carries **5** marks.
- (viii) There is no overall choice given in the question paper. However, an internal choice has been provided in few questions in all the sections except Section A.
- (ix) Kindly note that there is a separate question paper for Visually Impaired candidates.
- (x) Use of calculator is **not** allowed.

SECTION A

Questions no. **1 to 16** are Multiple Choice type Questions, carrying **1** mark each. $16 \times 1 = 16$

1. Which by-product is obtained in the manufacture of phenol from cumene ?
(A) Acetaldehyde (B) Acetone
(C) Anisole (D) *o*-Cresol
2. Which reagents are required for one step conversion of chlorobenzene to toluene ?
(A) $\text{NaNO}_2 + \text{HCl}$ (B) $\text{CH}_3\text{Cl}/\text{Anhydrous AlCl}_3$
(C) $\text{CH}_3\text{Cl}/\text{Fe}$, Dark (D) CH_3Cl , Na, Dry ether
3. Coordination number of Fe in $[\text{Fe}(\text{C}_2\text{O}_4)_3]^{3-}$ is :
(A) 6 (B) 3
(C) 4 (D) 5



4. Which of the following complexes shows geometrical isomerism ?
- (A) $[\text{Co}(\text{NH}_3)_6]^{3+}$
(B) $[\text{Co}(\text{NH}_3)_5\text{Cl}]^{2+}$
(C) $[\text{Co}(\text{NH}_3)_4\text{Cl}_2]^+$
(D) $[\text{Co}(\text{NH}_3)_5(\text{ONO})]^{2+}$
5. Which of the following does **not** show variable oxidation state ?
- (A) Sc (B) Mn
(C) Cr (D) Cu
6. The reagent that can be used to convert benzenediazonium chloride to benzene is :
- (A) Cu/HCl (B) H_2O
(C) $\text{CH}_3\text{CH}_2\text{OH}$ (D) CuCN
7. The gas evolved when methylamine reacts with HNO_2 is :
- (A) NO (B) N_2
(C) NH_3 (D) NO_2
8. A solution of acetone in chloroform :
- (A) obeys Raoult's law.
(B) forms azeotrope.
(C) shows a positive deviation from Raoult's law.
(D) shows a negative deviation from Raoult's law.



9. The freezing point of one molal KCl solution, assuming KCl to be completely dissociated in water, is : (K_f for water = $1.86 \text{ K kg mol}^{-1}$)
- (A) -3.72°C
(B) $+3.72^\circ\text{C}$
(C) -1.86°C
(D) $+2.72^\circ\text{C}$
10. A galvanic cell can behave like an electrolytic cell when
- (A) $E_{\text{cell}} = E_{\text{ext}}$ (B) $E_{\text{ext}} > E_{\text{cell}}$
(C) $E_{\text{cell}} > E_{\text{ext}}$ (D) $E_{\text{cell}} = 0$
11. The value of rate constant for a pseudo first order reaction :
- (A) depends only on temperature.
(B) depends on the concentration of reactants present in small amount.
(C) depends on the concentration of reactants present in large excess.
(D) is not dependent on the concentration of reactants.
12. Furanose ring of fructose is formed due to reaction between :
- (A) C_1 and C_5 (B) C_2 and C_5
(C) C_1 and C_4 (D) C_1 and C_2



For Questions number 13 to 16, two statements are given — one labelled as Assertion (A) and the other labelled as Reason (R). Select the correct answer to these questions from the codes (A), (B), (C) and (D) as given below.

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
- (B) Both Assertion (A) and Reason (R) are true, but Reason (R) is **not** the correct explanation of the Assertion (A).
- (C) Assertion (A) is true, but Reason (R) is false.
- (D) Assertion (A) is false, but Reason (R) is true.

13. *Assertion (A) :* Actinoids show wide range of oxidation states.

Reason (R) : This is due to comparable energies of 5f, 6d and 7s orbitals.

14. *Assertion (A) :* Order of the reaction can be zero or fractional.

Reason (R) : We cannot determine order from balanced chemical equation.

15. *Assertion (A) :* Boiling point of $(C_2H_5)_2NH$ is lower than that of $n-C_4H_9NH_2$.

Reason (R) : Hydrogen bonding is much more extensive in $n-C_4H_9NH_2$ as compared to $(C_2H_5)_2NH$.

16. *Assertion (A) :* Phenol is less acidic than alcohol.

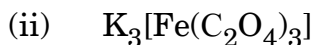
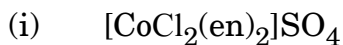
Reason (R) : Phenoxide ion is more stable than alkoxide ion.

SECTION B

17. State Raoult's law for a solution containing volatile components. Why is the vapour pressure of an aqueous solution of glucose lower than that of water ?



18. (a) Write IUPAC names of the following coordination compounds : 1+1=2



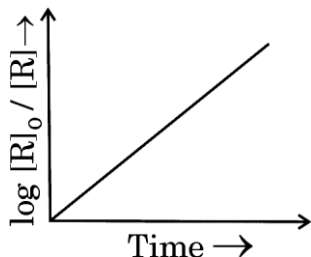
OR

(b) Differentiate between : 1+1=2

(i) Double salt and Complex compound

(ii) Didentate ligand and Ambidentate ligand

19. Observe the graph in the given figure and answer the following questions : 1+1=2



(a) Predict the order of reaction.

(b) What is the slope of the curve ?

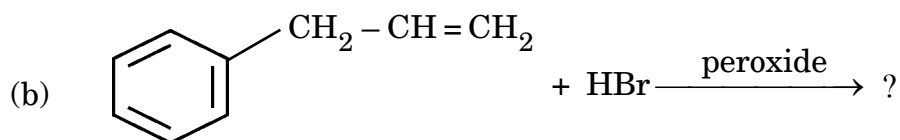
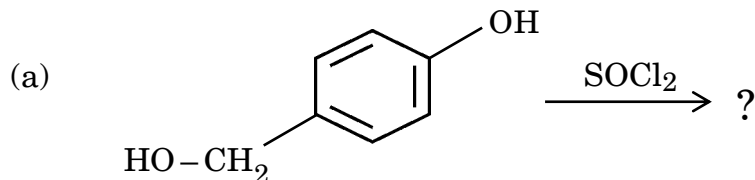
20. How do you explain the following ? 1+1=2

(a) Presence of an aldehydic group in glucose.

(b) Presence of five – OH groups in glucose.



21. Draw the structures of the major monohalo products in each of the following reactions : 1+1=2



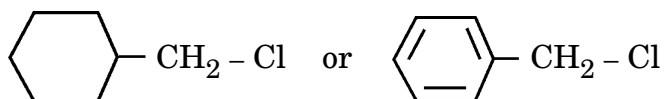
SECTION C

22. How can you obtain the following from aniline ? Give only chemical equation. 3×1=3

- (a) *p*-nitroaniline
- (b) Chlorobenzene
- (c) Phenol

23. A compound (A) with molecular formula C_4H_5N on reduction with DIBAL-H followed by hydrolysis, gives a compound (B). Compound (B) gives positive Tollens' test but does not give iodoform test. Compound (B) can also be obtained when ethanal is treated with dilute NaOH followed by heating. Identify (A) and (B). Write the reactions of (A) with DIBAL-H followed by hydrolysis. 3

24. Write any two differences between S_N1 and S_N2 reactions. Which of the following compounds would undergo S_N1 reaction faster and why ? 3





25. Vapour pressure of pure water at 298 K is 24.8 mm Hg. Calculate the lowering in vapour pressure of an aqueous solution which freezes at -0.3°C . (K_f of water = $1.86 \text{ K kg mol}^{-1}$) 3

26. (a) Answer the following about the complexes $3 \times 1 = 3$

$[\text{FeF}_6]^{3-}$ and $[\text{Fe}(\text{CN})_6]^{4-}$:

- Write the hybridization involved in each case.
 - Which of them is the outer orbital complex and which one is the inner orbital complex ?
 - Compare their magnetic behaviour.
- [Atomic number : Fe = 26]

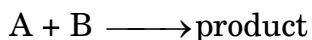
OR

- What happens to the colour of complex $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$ when heated gradually ?
 - Write the electronic configuration for d^5 ion if $\Delta_o < P$.
 - Write the hybridization and magnetic behaviour of the complex $[\text{Ni}(\text{CO})_4]$.
- [Atomic number : Ni = 28] $3 \times 1 = 3$

27. Give reasons for the following : $3 \times 1 = 3$

- The pH of aqueous NaCl increases when it is electrolysed.
- Unlike dry cell, mercury cell has a constant cell potential through its lifetime.
- Conductivity of solution decreases with dilution.

28. The rate of a reaction :



is given below as a function of different initial concentrations of A and B.

Experiment	[A] / mol L ⁻¹	[B] / mol L ⁻¹	Initial Rate/mol L ⁻¹ min ⁻¹
1	0.01	0.01	5×10^{-3}
2	0.02	0.01	1×10^{-2}
3	0.01	0.02	5×10^{-3}

Calculate the order of the reaction with respect to A and B. Determine the rate constant of the reaction. 3





SECTION D

The following questions are case-based questions. Read the case carefully and answer the questions that follow.

- 29.** The α -amino acids are the building blocks of proteins. All α -amino acids exist as zwitter ion due to which they show amphoteric behaviour. All amino acids are joined through peptide bond. Proteins are broadly classified as globular proteins and fibrous proteins. Globular proteins are water soluble, whereas fibrous proteins are not. The complete structure of protein is discussed at four different levels i.e. primary, secondary, tertiary and quaternary structures. Protein loses its biological activity in denatured form.
- (a) Define the following : 2
(i) Peptide linkage (ii) Denatured protein
- (b) Why do amino acids show amphoteric behaviour ? 1
- (c) (i) How can you differentiate between Fibrous protein and Globular protein ? 1
- OR**
- (c) (ii) Write the names of two different secondary structures of proteins. 1
- 30.** Alcohols undergo a number of reactions involving the cleavage of C – OH bond. However, phenols do not undergo reactions involving the cleavage of C – OH bond. Alcohols are weaker acids than water. Alcohols react with halogen acids to form the corresponding haloalkanes. Phenols are stronger acids than alcohols. A characteristic feature of phenols is that they undergo electrophilic substitution reactions such as halogenation, nitration, etc. Since – OH group is a strong activating group, phenol gives trisubstituted products during halogenation, nitration, etc.
- (a) What happens when phenol is treated with the following ? 2
(i) Br_2 water (ii) Conc. HNO_3
- (b) (i) Write the mechanism of alcohol reacting as nucleophile in a reaction with $\text{CH}_3^\oplus$. 1
- OR**
- (b) (ii) Why do phenols not undergo reactions involving cleavage of C – OH bond ? 1

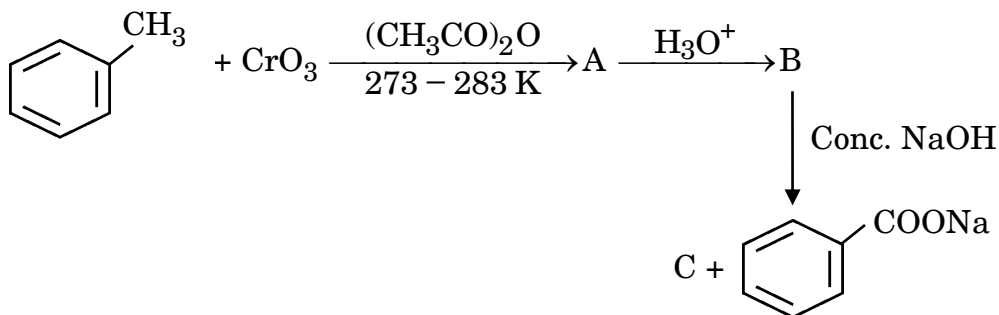


- (c) How can you distinguish between Butan-1-ol and 2-Methylpropan-2-ol by using HCl in the presence of anhydrous ZnCl_2 ?

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SECTION E

31. (a) (i) Identify A, B and C in the following reactions :

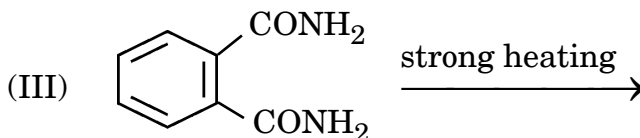
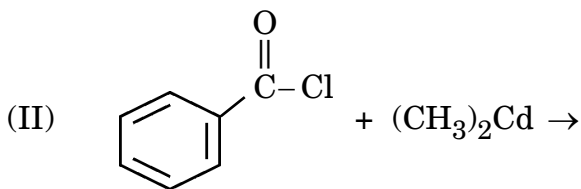


- (ii) Give reasons for the following :

- (I) Carboxylic acids do not give the characteristic reactions of carbonyl group.
 (II) Ethanoic acid is a stronger acid than ethanol. 3+2=5

OR

- (b) (i) Write the product(s) in the following reactions :



- (ii) Write the reaction involved in the following reactions :

- (I) Wolff-Kishner Reduction
 (II) Decarboxylation Reaction

3+2=5



32. (a) (i) Calculate E_{cell} of a galvanic cell in which the following reaction takes place at 25°C :
- $$\text{Zn(s)} + \text{Pb}^{2+}(0.02 \text{ M}) \longrightarrow \text{Zn}^{2+}(0.1 \text{ M}) + \text{Pb(s)}$$
- [Given : $E^\circ_{\text{Zn}^{2+}/\text{Zn}} = -0.76 \text{ V}$, $E^\circ_{\text{Pb}^{2+}/\text{Pb}} = -0.13 \text{ V}$;
 $\log 2 = 0.3010$, $\log 4 = 0.6021$, $\log 5 = 0.6990$].
- (ii) State Faraday's first law of electrolysis. How much electricity, in terms of Faraday, is required to reduce one mol of MnO_4^- to Mn^{2+} ion ? 3+2=5

OR

- (b) (i) The resistance of a conductivity cell containing 0.001 M KCl solution at 298 K is 1000 ohm . What is the cell constant if conductivity of 0.001 M KCl solution at 298 K is $0.125 \times 10^{-3} \text{ S cm}^{-1}$?
- (ii) Calculate the $E_{\text{Mg}^{2+}/\text{Mg}}$ potential for the following half cell at 25°C :
- $$\text{Mg}/\text{Mg}^{2+} (1 \times 10^{-4} \text{ M}); E^\circ_{\text{Mg}^{2+}/\text{Mg}} = +2.36 \text{ V}$$
- [Given : $\log 10 = 1$]
- (iii) What is the effect of temperature on the electrical conductance of metallic conductor ? 2+2+1=5

33. (a) (i) Account for the following :
- (I) Orange colour of $\text{Cr}_2\text{O}_7^{2-}$ ion changes to yellow when treated with an alkali.
- (II) Zn , Cd and Hg are non-transition elements.
- (III) E° value for $\text{Mn}^{3+}/\text{Mn}^{2+}$ couple is highly positive ($+1.57 \text{ V}$) as compared to $\text{Cr}^{3+}/\text{Cr}^{2+}$.
- (ii) What happens when :
- (I) Manganate ion undergoes disproportionation reaction in acidic medium ?
- (II) KMnO_4 is heated ? 3+2=5

OR



(b) Answer the following questions :

5×1=5

- (i) What is 'Misch metal' ? Give its one use.
- (ii) Write the formula of an oxoanion of chromium in which it shows the oxidation state equal to its group number.
- (iii) Why does Vanadium pentoxide (V_2O_5) act as a catalyst ?
- (iv) Why do transition elements have high enthalpies of atomisation ?
- (v) How do you prepare $Na_2Cr_2O_7$ from Na_2CrO_4 ?

