

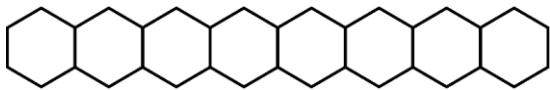


Series : W4YXZ

SET ~ 2

रोल नं.

Roll No.



प्रश्न-पत्र कोड

Q.P. Code

56/4/2

परीक्षार्थी प्रश्न-पत्र कोड को उत्तर-पुस्तिका के मुख-पृष्ठ पर अवश्य लिखें।

Candidates must write the Q.P. Code on the title page of the answer-book.

(I) कृपया जाँच कर लें कि इस प्रश्न-पत्र में मुद्रित पृष्ठ **23** हैं।

Please check that this question paper contains **23** printed pages.

(II) प्रश्न-पत्र में दाहिने हाथ की ओर दिए गए प्रश्न-पत्र कोड को परीक्षार्थी उत्तर-पुस्तिका के मुख-पृष्ठ पर लिखें।

Q.P. Code given on the right hand side of the question paper should be written on the title page of the answer-book by the candidate.

(III) कृपया जाँच कर लें कि इस प्रश्न-पत्र में **33** प्रश्न हैं।

Please check that this question paper contains **33** questions.

(IV) कृपया प्रश्न का उत्तर लिखना शुरू करने से पहले, उत्तर-पुस्तिका में यथा स्थान पर प्रश्न का क्रमांक अवश्य लिखें।

Please write down the Serial Number of the question in the answer-book at the given place before attempting it.

(V) इस प्रश्न-पत्र को पढ़ने के लिए 15 मिनट का समय दिया गया है। प्रश्न-पत्र का वितरण पूर्वाह्न में 10.15 बजे किया जाएगा। 10.15 बजे से 10.30 बजे तक परीक्षार्थी केवल प्रश्न-पत्र को पढ़ेंगे और इस अवधि के दौरान वे उत्तर-पुस्तिका पर कोई उत्तर नहीं लिखेंगे।

15 minute time has been allotted to read this question paper. The question paper will be distributed at 10.15 a.m. From 10.15 a.m. to 10.30 a.m., the candidates will read the question paper only and will not

write any answer on the answer-book during this period.



रसायन विज्ञान (सैद्धान्तिक)

CHEMISTRY (Theory)



निर्धारित समय : 3 घण्टे

Time allowed : 3 hours

अधिकतम अंक : 70

Maximum Marks : 70

56/4/2

1

P.T.O.



General Instructions :

Read the following instructions carefully and follow them :

- (i) This question paper contains **33** questions. **All** questions are **compulsory**.
- (ii) This question paper is divided into **five** sections – **Section A, B, C, D and E**.
- (iii) **Section A** – questions number **1 to 16** are multiple choice type questions. Each question carries **1** mark.
- (iv) **Section B** – questions number **17 to 21** are very short answer type questions. Each question carries **2** marks.
- (v) **Section C** – questions number **22 to 28** are short answer type questions. Each question carries **3** marks.
- (vi) **Section D** – questions number **29 and 30** are case-based questions. Each question carries **4** marks.
- (vii) **Section E** – questions number **31 to 33** are long answer type questions. Each question carries **5** marks.
- (viii) There is no overall choice given in the question paper. However, an internal choice has been provided in few questions in all the sections except Section A.
- (ix) Kindly note that there is a separate question paper for Visually Impaired candidates.
- (x) Use of calculator is **not** allowed.

SECTION A

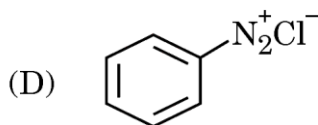
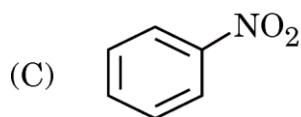
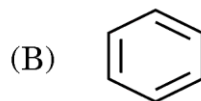
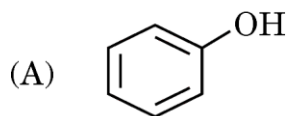
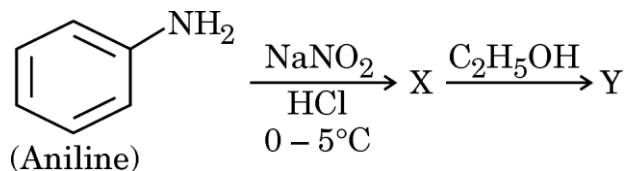
Questions no. **1 to 16** are Multiple Choice type Questions, carrying **1** mark each. $16 \times 1 = 16$

1. Alkenes are formed by heating alcohols with conc. H_2SO_4 . The first step in the reaction is :
 - (A) formation of carbocation
 - (B) formation of ester
 - (C) protonation of alcohol molecule
 - (D) elimination of water
2. Polyhalogen compounds have wide application in industries and agriculture. DDT is also a very important polyhalogen compound. It is a :
 - (A) greenhouse gas
 - (B) fertilizer
 - (C) biodegradable insecticide
 - (D) non-biodegradable insecticide
3. The product of the oxidation of I^- with MnO_4^- in alkaline medium is :

(A) IO_4^-	(B) I_2
(C) IO^-	(D) IO_3^-



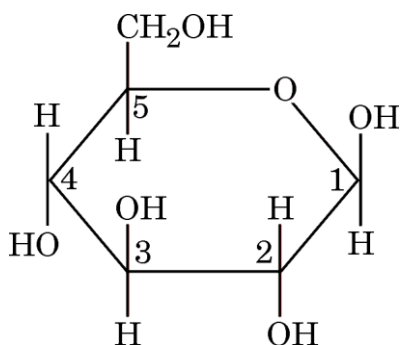
4. In the given reaction sequence, the structure of Y would be :



5. Out of 2-Bromobutane, 1-Bromobutane, 2-Bromopropane and 1-Bromopropane, the molecule which is chiral in nature is :

- (A) 2-Bromobutane
(B) 1-Bromobutane
(C) 2-Bromopropane
(D) 1-Bromopropane

6. In the Haworth structure of the following carbohydrate, various carbon atoms have been numbered. The anomeric carbon is numbered as :



- (A) 1
(B) 2
(C) 3
(D) 5



7. The value of Henry's constant K_H is :
- (A) greater for gases with higher solubility
 - (B) greater for gases with lower solubility
 - (C) constant for all gases
 - (D) not related to the solubility of gases
8. Out of the following statements, the **incorrect** statement is :
- (A) La is actually an element of transition series.
 - (B) Zr and Hf have almost identical atomic radii because of lanthanoid contraction.
 - (C) Ionic radius decreases from La^{3+} to Lu^{3+} ion.
 - (D) Lanthanoids are radioactive in nature.
9. In an electrochemical cell, the following reaction takes place :
- $$2\text{Ag}^+ (\text{aq}) + \text{Mg} (\text{s}) \rightarrow 2\text{Ag} (\text{s}) + \text{Mg}^{2+} (\text{aq})$$
- $$E^\circ_{\text{cell}} = 2.96 \text{ V}$$
- As the reaction progresses, what will happen to the overall voltage of the cell ?
- (A) Voltage will remain constant.
 - (B) It will decrease as $[\text{Mg}^{2+}]$ increases.
 - (C) It will increase as $[\text{Ag}^+]$ increases.
 - (D) It will increase as $[\text{Mg}^{2+}]$ increases.
10. Out of Ti^{3+} , Cr^{3+} , Mn^{2+} and Ni^{2+} ions, the one which has the highest magnetic moment is :
- (A) Ti^{3+}
 - (B) Cr^{3+}
 - (C) Mn^{2+}
 - (D) Ni^{2+}

[Atomic number : Ti = 22, Cr = 24, Mn = 25, Ni = 28]



11. Hoffmann Bromamide degradation reaction is given by :
- (A) CH_3NO_2 (B) CH_3NH_2
(C) CH_3CONH_2 (D) $\text{CH}_3\text{CH}_2\text{NH}_2$
12. What amount of electric charge is required for the oxidation of 1 mole of H_2O to O_2 ?
- (A) 1 F (B) 2 F
(C) 3 F (D) 4 F

For Questions number 13 to 16, two statements are given — one labelled as Assertion (A) and the other labelled as Reason (R). Select the correct answer to these questions from the codes (A), (B), (C) and (D) as given below.

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
(B) Both Assertion (A) and Reason (R) are true, but Reason (R) is **not** the correct explanation of the Assertion (A).
(C) Assertion (A) is true, but Reason (R) is false.
(D) Assertion (A) is false, but Reason (R) is true.
13. *Assertion (A) :* Cuprous salts are diamagnetic.
Reason (R) : In cuprous ion, 3d-orbitals are partially filled.
14. *Assertion (A) :* Acetanilide is more basic than aniline.
Reason (R) : Acetylation of aniline results in decrease of electron density on nitrogen.
15. *Assertion (A) :* Electrolysis of aqueous NaCl gives H_2 at cathode and Cl_2 at anode.
Reason (R) : Chlorine has higher oxidation potential than H_2O .
16. *Assertion (A) :* n-Butyl chloride has higher boiling point than n-Butyl bromide.
Reason (R) : C – Cl bond is more polar than C – Br bond.



SECTION B

17. What is meant by essential amino acids ? Why are amino acids amphoteric in nature ? 2

18. (a) Reactant 'A' underwent a decomposition reaction. The concentration of 'A' was measured periodically and recorded in the table given below :

Time/Hours	[A]/M
0	0.40
1	0.20
2	0.10
3	0.05

Based on the above data, predict the order of the reaction and write the expression for the rate law. 2

OR

(b) The reaction between H_2 (g) and I_2 (g) was carried out in a sealed isothermal container. The rate law for the reaction was found to be :

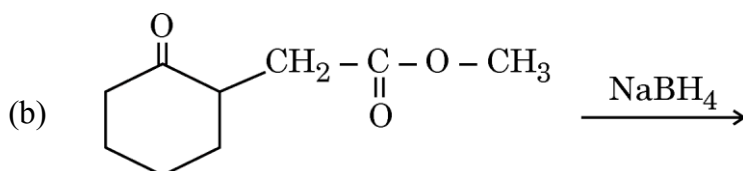
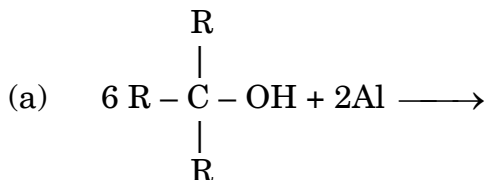
$$\text{Rate} = k[\text{H}_2] [\text{I}_2]$$

If 1 mole of H_2 (g) was added to the reaction chamber and the temperature was kept constant, then predict the change in rate of the reaction and the rate constant. 2

19. $\text{PtCl}_4 \cdot 2\text{KCl}$ doesn't give precipitate of AgCl with AgNO_3 solution. Write the structural formula and IUPAC name of the complex. 2

20. Define fuel cell. Give two advantages of fuel cell over ordinary cell. 2

21. Write the structures of the main products of the following reactions : 2





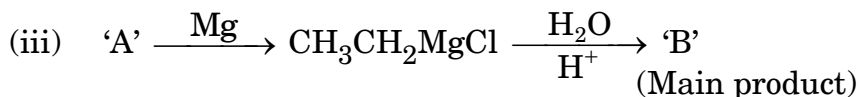
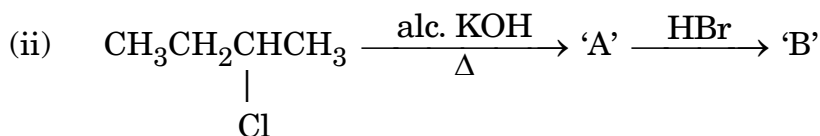
SECTION C

22. (a) Can sodium methoxide and t-butyl bromide be used for the preparation of t-butyl methyl ether. Give suitable reason. Justify your answer by suggesting the appropriate starting material if required for the preparation of t-butyl methyl ether. 2
- (b) Give the IUPAC name of above mentioned ether. 1
23. Arrange the following compounds as asked : 3
- (a) in increasing order of pK_b values
 $C_2H_5NH_2$, $(C_2H_5)_2NH$, $C_6H_5NHCH_3$, $C_6H_5NH_2$
- (b) in decreasing order of boiling point
 C_2H_5OH , $C_2H_5NH_2$, $(CH_3)_2NH$
- (c) in decreasing order of solubility in water
 $C_6H_5NH_2$, $(C_2H_5)_2NH$, $C_2H_5NH_2$
24. Calculate the cell voltage of the voltaic cell which is set up by joining the following half-cells at $25^\circ C$: 3
- Al/Al^{3+} (0.001 M) and Ni/Ni^{2+} (0.001 M)
- Given : $E^\circ_{Ni^{2+}/Ni} = -0.25 V$
- $E^\circ_{Al^{3+}/Al} = -1.66 V$
25. (a) Account for the following : 3
- (i) Allyl chloride is hydrolysed more readily than n-propyl chloride.
- (ii) Isocyanides are formed when alkyl halides are treated with silver cyanide.
- (iii) Methyl chloride reacts faster with $\bar{O}H$ ion in S_N2 reaction than t-butyl chloride.

OR



- (b) Complete the following reactions by writing the structural formulae of 'A' and 'B' : 3



26. Give explanation for each of the following observations : 3

- (a) With the same d-orbital configuration (d^4), Mn^{3+} ion is an oxidising agent whereas Cr^{2+} ion is a reducing agent.
- (b) Actinoid contraction is greater from element to element than that among lanthanoids.
- (c) Transition metals form large number of interstitial compounds with H, B, C and N.

27. An aqueous solution of NaOH was made and its molar mass from the measurement of osmotic pressure at 27°C was found to be 25 g mol^{-1} . Calculate the percentage dissociation of NaOH in this solution. 3

[Atomic mass : Na = 23 u, O = 16 u, H = 1 u]

28. A compound 'A' with molecular formula $\text{C}_4\text{H}_8\text{O}$ gives positive 2,4-DNP test. It gives yellow precipitate of compound 'B' on treatment with sodium hypoiodite. Compound 'A' does not react with Tollen's or Fehling's reagent; on drastic oxidation with KMnO_4 , it forms a carboxylic acid 'C'. Elucidate the structures of 'A', 'B' and 'C'. Also give their IUPAC names. 3



SECTION D

The following questions are case-based questions. Read the case carefully and answer the questions that follow.

- 29.** Ribose and 2-deoxyribose have an important role in biology. Among the most important derivatives are those with phosphate groups attached at the 5 position. Mono-, di- and tri-phosphate forms are important, as well as 3-5 cyclic monophosphates. Purines and pyrimidines form an important class of compounds with ribose and deoxyribose. When these purine and pyrimidine derivatives are coupled to a ribose sugar, they are called nucleosides.

Answer the following questions :

- (a) What products would be formed when DNA is hydrolysed ? How is DNA different from RNA with reference to a structure ? 2
- (b) Differentiate between nucleotide and nucleoside. 1
- (c) (i) Mention two important functions of nucleic acid. 1

OR

- (c) (ii) Name the linkage which joins two nucleotides. Name the base that is found in nucleotide of RNA but not in DNA. 1

- 30.** According to the generally accepted definition of the ideal solution there are equal interaction forces acting between molecules belonging to the same or different species. (This is equivalent to the statement that the activity of the components equals the concentration.) Strictly speaking, this condition is fulfilled only in exceptional cases for mixtures (optical isomers, isotopic mixtures of an element, hydrocarbon mixtures). It is still usual to talk about ideal solutions as limiting cases in reality since very dilute solutions behave ideally with respect to the solvent. This view is further supported by the fact that Raoult's law empirically found for describing the behaviour of the solvent in dilute solutions can be deduced thermodynamically via the assumption of ideal behaviour of the solvent.

Answer the following questions :

- (a) Give one example of miscible liquid pair which shows negative deviation from Raoult's law. What is the reason for such deviation ? 2
- (b) (i) State Raoult's law for a solution containing volatile components. 1

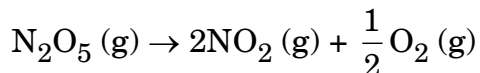
OR

- (b) (ii) Raoult's law is a special case of Henry's law. Comment. 1
- (c) Write two characteristics of an ideal solution. 1



SECTION E

31. (a) (i) The initial concentration of N_2O_5 in the first order reaction :



was $1.2 \times 10^{-2} \text{ mol L}^{-1}$. The concentration of N_2O_5 after 60 minutes was $0.2 \times 10^{-2} \text{ mol L}^{-1}$. Calculate the rate constant of the reaction at 318 K.

3

$$[\log 6 = 0.778]$$

- (ii) Account for the following :

(I) We cannot determine the order of a reaction by taking into consideration the balanced chemical equation.

1

(II) A bimolecular reaction may become kinetically of first order under a specified condition.

1

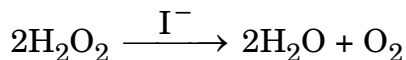
OR

- (b) (i) The rate of the chemical reaction doubles for an increase of 10 K in absolute temperature from 298 K. Calculate activation energy (E_a).

3

$$[2.303 R = 19.15 \text{ JK}^{-1} \text{ mol}^{-1}, \log 2 = 0.3]$$

- (ii) For a reaction :



the proposed mechanism is as given below :

(I) $\text{H}_2\text{O}_2 + \text{I}^- \longrightarrow \text{H}_2\text{O} + \text{IO}^-$ (slow)

(II) $\text{H}_2\text{O}_2 + \text{IO}^- \longrightarrow \text{H}_2\text{O} + \text{I}^- + \text{O}_2$ (fast)

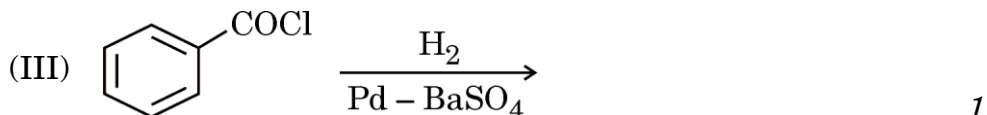
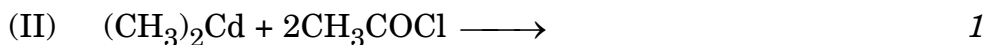
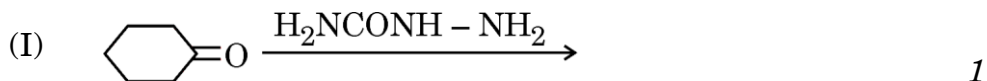
(1) Write rate law for the reaction.

(2) Write the overall order and molecularity of the reaction.

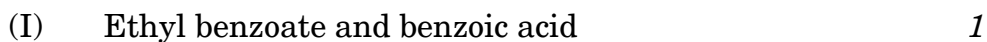
2



32. (a) (i) Complete the following reactions by writing the structure of the main products :

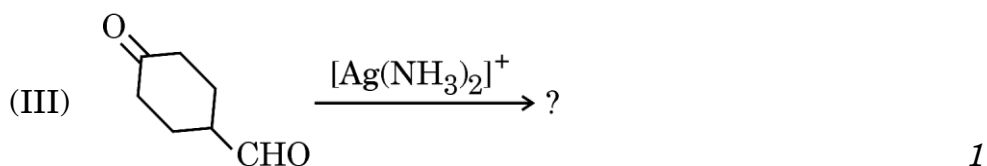
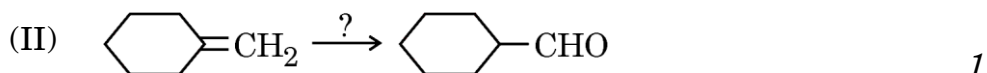
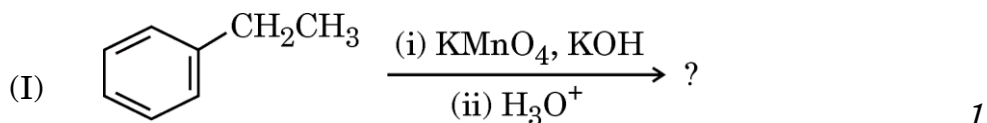


- (ii) Give simple chemical test to distinguish between the following pairs of compounds :

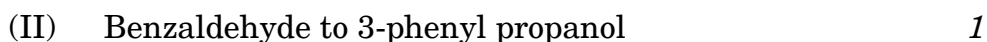


OR

- (b) (i) Complete each synthesis by giving missing starting material, reagent or products :



- (ii) Carry out the following conversions :





- 33.** (a) (i) Give reasons :
- (I) $[\text{Ni}(\text{CO})_4]$ is diamagnetic whereas $[\text{NiCl}_4]^{2-}$ is paramagnetic. [Atomic number : Ni = 28] 1
- (II) CO is a stronger complexing agent than NH_3 . 1
- (III) The trans isomer of complex $[\text{Co}(\text{en})_2\text{Cl}_2]^+$ is optically inactive. 1
- (ii) Using Crystal Field theory, write the number of unpaired electrons in octahedral complexes of Fe^{3+} in the presence of : 2
- (I) Strong field ligand
- (II) Weak field ligand
- [Atomic number : Fe = 26]
- OR**
- (b) (i) Name the type of isomerism exhibited by the following compounds. Also draw their corresponding isomers.
- (I) $[\text{Co}(\text{NH}_3)_6] [\text{Cr}(\text{CN})_6]$ 1
- (II) $[\text{Co}(\text{en})_3]^{3+}$ 1
- (III) $[\text{Co}(\text{NH}_3)_3(\text{NO}_2)_3]$ 1
- (ii) Differentiate between weak field and strong field ligands. How does the strength of the ligand influence the spin of the complex ? 2

