

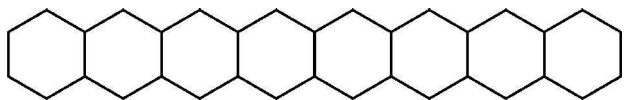


Series : ZYWX1

SET ~ 2



रोल नं.
Roll No.



प्रश्न-पत्र कोड
Q.P. Code **56/1/2**

परीक्षार्थी प्रश्न-पत्र कोड को उत्तर-पुस्तिका के मुख-पृष्ठ पर अवश्य लिखें।

Candidates must write the Q.P. Code on the title page of the answer-book.

नोट / NOTE



- (I) कृपया जाँच कर लें कि इस प्रश्न-पत्र में मुद्रित पृष्ठ 23 हैं।
Please check that this question paper contains 23 printed pages.
- (II) कृपया जाँच कर लें कि इस प्रश्न-पत्र में 33 प्रश्न हैं।
Please check that this question paper contains 33 questions.
- (III) प्रश्न-पत्र में दाहिने हाथ की ओर दिए गए प्रश्न-पत्र कोड को परीक्षार्थी उत्तर-पुस्तिका के मुख-पृष्ठ पर लिखें।
Q.P. Code given on the right hand side of the question paper should be written on the title page of the answer-book by the candidate.
- (IV) कृपया प्रश्न का उत्तर लिखना शुरू करने से पहले, उत्तर-पुस्तिका में यथा स्थान पर प्रश्न का क्रमांक अवश्य लिखें।
Please write down the serial number of the question in the answer-book at the given place before attempting it.
- (V) इस प्रश्न-पत्र को पढ़ने के लिए 15 मिनट का समय दिया गया है। प्रश्न-पत्र का वितरण पूर्वाह्न में 10.15 बजे किया जाएगा। 10.15 बजे से 10.30 बजे तक परीक्षार्थी केवल प्रश्न-पत्र को पढ़ेंगे और इस अवधि के दौरान वे उत्तर-पुस्तिका पर कोई उत्तर नहीं लिखेंगे।
15 minute time has been allotted to read this question paper. The question paper will be distributed at 10.15 a.m. From 10.15 a.m. to 10.30 a.m., the candidates will read the question paper only and will not write any answer on the answer-book during this period.



रसायन विज्ञान (सैद्धांतिक)
CHEMISTRY (Theory)



निर्धारित समय : 3 घण्टे

Time allowed : 3 hours

अधिकतम अंक : 70

Maximum Marks : 70

56/1/2

730-2

1

P.T.O.



General Instructions :

Read the following instructions carefully and follow them :

- (i) *This question paper contains 33 questions. All questions are compulsory.*
- (ii) *This question paper is divided into five sections – Section A, B, C, D and E.*
- (iii) *Section A – questions number 1 to 16 are multiple choice type questions. Each question carries 1 mark.*
- (iv) *Section B – questions number 17 to 21 are very short answer type questions. Each question carries 2 marks.*
- (v) *Section C – questions number 22 to 28 are short answer type questions. Each question carries 3 marks.*
- (vi) *Section D – questions number 29 and 30 are case-based questions. Each question carries 4 marks.*
- (vii) *Section E – questions number 31 to 33 are long answer type questions. Each question carries 5 marks.*
- (viii) *There is no overall choice given in the question paper. However, an internal choice has been provided in few questions in all the sections except Section –A.*
- (ix) *Kindly note that there is a separate question paper for Visually Impaired candidates.*
- (x) *Use of calculator is **NOT** allowed.*

You may use the following values of physical constants wherever necessary :

$$c = 3 \times 10^8 \text{ m/s}$$

$$h = 6.63 \times 10^{-34} \text{ Js}$$

$$e = 1.6 \times 10^{-19} \text{ C}$$

$$\mu_0 = 4\pi \times 10^{-7} \text{ T m A}^{-1}$$

$$\epsilon_0 = 8.854 \times 10^{-12} \text{ C}^2 \text{ N}^{-1} \text{ m}^{-2}$$

$$\frac{1}{4\pi\epsilon_0} = 9 \times 10^9 \text{ N m}^2 \text{ C}^{-2}$$

$$\text{Mass of electron (m}_e\text{)} = 9.1 \times 10^{-31} \text{ kg.}$$

$$\text{Mass of neutron} = 1.675 \times 10^{-27} \text{ kg.}$$

$$\text{Mass of proton} = 1.673 \times 10^{-27} \text{ kg.}$$

$$\text{Avogadro's number} = 6.023 \times 10^{23} \text{ per gram mole}$$

$$\text{Boltzmann's constant} = 1.38 \times 10^{-23} \text{ JK}^{-1}$$



SECTION – A**(16 × 1 = 16)**

Questions No. 1 to 16 are Multiple Choice type questions carrying 1 mark each.

For questions number 1 to 4, two statements are given – one labelled as Assertion (A) and the other labelled as Reason (R). Select the correct answer to these questions from the codes (A), (B), (C) and (D) as given below :

- (A) Both Assertion (A) and Reason (R) are true and Reason (R) is the correct explanation of the Assertion (A).
- (B) Both Assertion (A) and Reason (R) are true, but Reason (R) is not the correct explanation of the Assertion (A).
- (C) Assertion (A) is true, but Reason (R) is false.
- (D) Assertion (A) is false, but Reason (R) is true.

1. **Assertion (A)** : All naturally occurring α -amino acids except glycine are optically active.

Reason (R) : Most naturally occurring amino acids have L-configuration.

2. **Assertion (A)** : The boiling point of ethanol is higher than that of methoxymethane.

Reason (R) : There is intramolecular hydrogen bonding in ethanol.

3. **Assertion (A)** : The boiling points of alkyl halides decrease in the order : $RI > RBr > RCl > RF$.

Reason (R) : The boiling points of alkyl chlorides, bromides and iodides are considerably higher than that of the hydrocarbon of comparable molecular mass.

4. **Assertion (A)** : $[Cr(H_2O)_6]Cl_2$ and $[Fe(H_2O)_6]Cl_2$ are examples of homoleptic complexes.

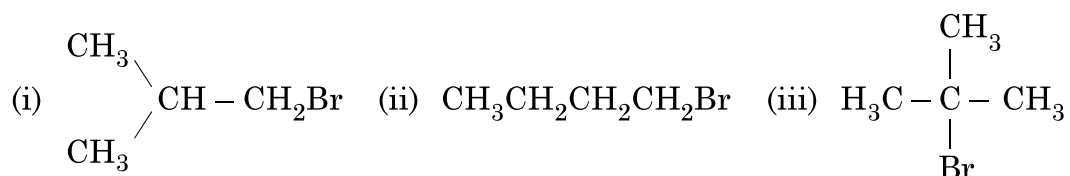
Reason (R) : All the ligands attached to the metal are the same.



5. An unripe mango placed in a concentrated salt solution to prepare pickle, shrivels because _____.
(A) it gains water due to osmosis
(B) it loses water due to reverse osmosis
(C) it gains water due to reverse osmosis
(D) it loses water due to osmosis
6. Which of the following statements is not true about glucose ?
(A) It is an aldohexose.
(B) On heating with HI it forms n-hexane.
(C) It exists in furanose form.
(D) It does not give Schiff's test.
7. The best reagent for converting propanamide into propanamine is _____.
(A) excess H_2
(B) Br_2 in aqueous NaOH
(C) iodine in the presence of red phosphorus
(D) $LiAlH_4$ in ether
8. The acid formed when propyl magnesium bromide is treated with CO_2 followed by acid hydrolysis is :
(A) C_3H_7COOH
(B) C_2H_5COOH
(C) CH_3COOH
(D) C_3H_7OH
9. Which is the correct order of acid strength from the following ?
(A) $C_6H_5OH > H_2O > ROH$
(B) $C_6H_5OH > ROH > H_2O$
(C) $ROH > C_6H_5OH > H_2O$
(D) $H_2O > C_6H_5OH > ROH$



10. Alkyl halides undergoing nucleophilic bimolecular substitution reaction involve
- (A) retention of configuration
(B) formation of racemic mixture
(C) inversion of configuration
(D) formation of carbocation
11. Arrange the following compounds in increasing order of their boiling points :



The correct order is

- (A) (ii) < (i) < (iii) (B) (i) < (ii) < (iii)
(C) (iii) < (i) < (ii) (D) (iii) < (ii) < (i)
12. The correct IUPAC name of $[\text{Pt}(\text{NH}_3)_2\text{Cl}_2]^{2+}$ is
- (A) Diamminedichloridoplatinum (II)
(B) Diamminedichloridoplatinum (IV)
(C) Diamminedichloridoplatinum (O)
(D) Diamminedichloridoplatinate (IV)

13. Acidified KMnO_4 oxidises sulphite to

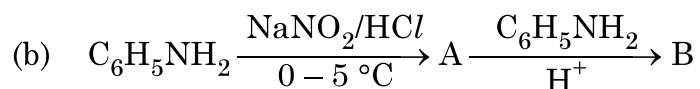
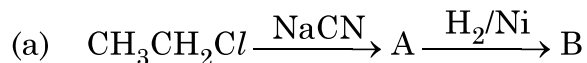
- (A) $\text{S}_2\text{O}_3^{2-}$ (B) $\text{S}_2\text{O}_8^{2-}$
(C) $\text{SO}_2(\text{g})$ (D) SO_4^{2-}



14. The magnetic moment is associated with its spin angular momentum and orbital angular momentum. Spin only magnetic moment value of Cr^{3+} ion (Atomic no. : Cr = 24) is _____.
(A) 2.87 B.M. (B) 3.87 B.M.
(C) 3.47 B.M. (D) 3.57 B.M.
15. Standard electrode potential for $\text{Sn}^{4+}/\text{Sn}^{2+}$ couple is +0.15 V and that for the Cr^{3+}/Cr couple is -0.74 V. The two couples in their standard states are connected to make a cell. The cell potential will be
(A) +1.19 V (B) +0.89 V
(C) +0.18 V (D) +1.83 V
16. In case of association, abnormal molar mass of solute will
(A) increase (B) decrease
(C) remain same (D) first increase and then decrease

SECTION – B

17. Identify A and B in each of the following reaction sequence : (1 + 1 = 2)



18. When FeCr_2O_4 is fused with Na_2CO_3 in the presence of air it gives a yellow solution of compound (A). Compound (A) on acidification gives compound (B). Compound (B) on reaction with KCl forms an orange coloured (C). An acidified solution of compound (C) oxidises iodide to (D). Identify (A), (B), (C) and (D). 2

19. Would you expect benzaldehyde to be more reactive or less reactive in nucleophilic addition reactions than propanal ? Justify your answer. 2



20. (A) Give reasons : (1 + 1 = 2)

- (a) Cooking is faster in pressure cooker than in an open pan.
- (b) On mixing liquid X and liquid Y, volume of the resulting solution decreases. What type of deviation from Raoult's law is shown by the resulting solution ? What change in temperature would you observe after mixing liquids X and Y ?

OR

(B) Define Azeotrope. What type of Azeotrope is formed by negative deviation from Raoult's law ? Give an example. 2

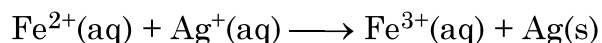
21. Give reasons for the following : 2

- (a) The melting points of α -amino acids are generally higher than that of the corresponding carboxylic acids.
- (b) Amino acids show amphoteric behaviour.

SECTION - C

22. A solution containing 15 g urea (molar mass = 60 g mol⁻¹) per litre of solution in water has the same osmotic pressure (isotonic) as a solution of glucose (molar mass = 180 g mol⁻¹) in water. Calculate the mass of glucose present in one litre of its solution. 3

23. Calculate $\Delta_r G^\circ$ and $\log K_C$ of the reaction : 3



Given $E^\circ_{\text{Ag}^+/\text{Ag}} = 0.80 \text{ V}$, $E^\circ_{\text{Fe}^{3+}/\text{Fe}^{2+}} = 0.77 \text{ V}$

[$R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$, $F = 96500 \text{ C mol}^{-1}$]



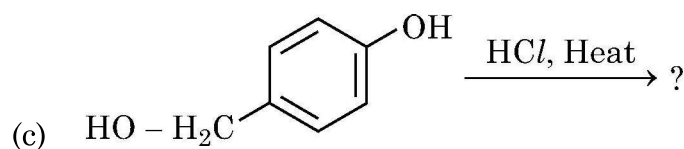
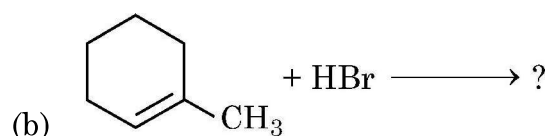
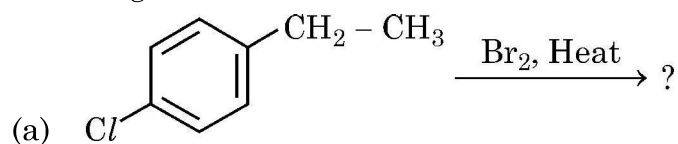
24. (a) Arrange the following in decreasing order of pK_b : (1 + 2 = 3)

Aniline, p-nitroaniline, p-methylaniline

- (b) Account for the following :

- (i) Diazonium salts of aromatic amines are more stable than those of aliphatic amines.
(ii) Methylamine in water reacts with $FeCl_3$ to precipitate hydrated ferric oxide.

25. (A) Draw the structure of the major monohalo product for each of the following reaction : (1 + 1 + 1 = 3)



OR

- (B) How do you convert : (3 × 1 = 3)

- (a) Chlorobenzene to biphenyl
(b) Propene to 1-Iodopropane
(c) 2-bromobutane to but-2-ene.

26. The elements of 3d transition series are given as : (1 + 1 + 1 = 3)

Sc, Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn

Answer the following :

- (a) Copper has exceptionally positive $E_{M^{2+}/M}^\circ$ value, why ?
(b) Which element is a strong reducing agent in +2 oxidation state and why ?
(c) Zn^{2+} salts are colourless. Why ?



-
27. A certain reaction is 50% complete in 20 minutes at 300 K and the same reaction is 50% complete in 5 minutes at 350 K. Calculate the activation energy if it is a first order reaction. 3
[$R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$; $\log 4 = 0.602$]
28. (a) Write the reaction when D-glucose reacts with the following : 3
(i) NH_2OH
(ii) Acetic anhydride
(b) Why vitamin C cannot be stored in our body ?

SECTION – D

The following questions are case based questions. Read the passage carefully and answer the questions that follow.

29. Phenols undergo electrophilic substitution reactions readily due to the strong activating effect of OH group attached to the benzene ring. Since, the OH group increases the electron density more to o- and p- positions therefore OH group is ortho, para-directing. Reimer-Tiemann reaction is one of the examples of aldehyde group being introduced on the aromatic ring of phenol, ortho to the hydroxyl group. This is a general method used for the ortho-formylation of phenols.

Answer the following questions :

- (a) What happens when phenol reacts with 2
(i) Br_2/CS_2
(ii) Conc. HNO_3
- (b) Why phenol does not undergo protonation readily ? 1
- (c) Which is a stronger acid – phenol or cresol ? Give reason. 1

OR

- (c) Write the IUPAC name of the product formed in the Reimer-Tiemann reaction. 1



30. The rate of a chemical reaction is expressed either in terms of decrease in the concentration of reactants or increase in the concentration of a product per unit time. Rate of the reaction depends upon the nature of reactants, concentration of reactants, temperature, presence of catalyst, surface area of the reactants and presence of light. Rate of reaction is directly related to the concentration of reactant. Rate law states that the rate of reaction depends upon the concentration terms on which the rate of reaction actually depends, as observed experimentally. The sum of powers of the concentration of the reactants in the Rate law expression is called order of reaction while the number of reacting species taking part in an elementary reaction which must collide simultaneously in order to bring about a chemical reaction is called molecularity of the reaction.

Answer the following questions :

- (a) (i) What is a rate determining step ? (1 + 1 = 2)
(ii) Define complex reaction.
(b) What is the effect of temperature on the rate constant of a reaction ? 1

OR

- (b) Why is molecularity applicable only for elementary reactions whereas order is applicable for elementary as well as complex reactions ? 1
(c) The conversion of molecule X to Y follows second order kinetics. If concentration of X is increased 3 times, how will it affect the rate of formation of Y ? 1

SECTION – E

31. (A) (a) Carry out the following conversions : (2 + 3 = 5)
(i) Ethanal to But-2-enal
(ii) Propanoic acid to ethane
(b) An alkene A with molecular formula C_5H_{10} on ozonolysis gives a mixture of two compounds B and C. Compound B gives positive Fehling test and also reacts with iodine and NaOH solution. Compound C does not give Fehling solution test but forms iodoform. Identify the compounds A, B and C.

OR



31. (B) An organic compound (A) (molecular formula $C_8H_{16}O_2$) was hydrolysed with dilute sulphuric acid to get a carboxylic acid (B) and an alcohol (C). Oxidation of (C) with chromic acid produced (B). (C) on dehydration gives But-1-ene. Identify (A), (B) and (C) and write chemical equations for the reactions involved. 5

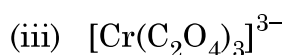
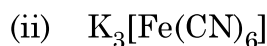
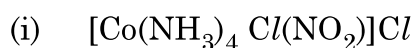
32. (A) For the complex $[Fe(en)_2Cl_2] Cl$, identify : 5

- (a) the oxidation number of iron.
- (b) the hybridization and the shape of the complex.
- (c) the magnetic behaviour of the complex
- (d) whether there is an optical isomer of the complex ? If so draw its structure.
- (e) IUPAC name of the complex.

(At. no. of Fe = 26)

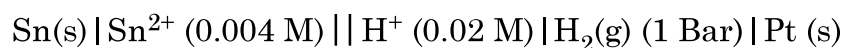
OR

32. (B) (a) Using IUPAC norms write the names of the following : (3 + 2 = 5)



- (b) What is crystal field splitting energy ? Why low spin tetrahedral complexes are not formed ?

33. (A) (a) Write the cell reaction and calculate the e.m.f. of the following cell at 298 K : (3 + 2 = 5)



(Given : $E^\circ_{Sn^{2+}/Sn} = -0.14 V$, $E^\circ_{H^+ | H_2(g), Pt} = 0.00V$)



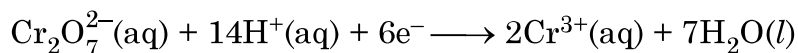
(b) Account for the following ;

- (i) On the basis of E° values, O_2 gas should be liberated at anode but it is Cl_2 gas which is liberated in the electrolysis of aqueous $NaCl$.
- (ii) Conductivity of CH_3COOH decreases on dilution.

OR

33. (B) (a) Write the anode and cathode reactions and the overall cell reaction occurring in a lead storage battery during its use. **(2 + 3 = 5)**
- (b) Calculate the potential for half-cell containing 0.01 M $K_2Cr_2O_7(aq)$, 0.01 M $Cr^{3+}(aq)$ and 1.0×10^{-4} M $H^+(aq)$.

The half cell reaction is



and the standard electrode potential is given as $E^\circ = 1.33$ V.

[Given : $\log 10 = 1$]

