

Series HFG1E/2



SET-2

प्रश्न-पत्र कोड
Q.P. Code 56/2/2

रोल नं.

Roll No.

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परीक्षार्थी प्रश्न-पत्र कोड को उत्तर-पुस्तिका के मुख-पृष्ठ पर अवश्य लिखें।

Candidates must write the Q.P. Code on the title page of the answer-book.

रसायन विज्ञान (सैद्धांतिक)

CHEMISTRY (Theory)

निर्धारित समय : 3 घण्टे

Time allowed : 3 hours

अधिकतम अंक : 70

Maximum Marks : 70

- कृपया जाँच कर लें कि इस प्रश्न-पत्र में मुद्रित पृष्ठ 23 हैं।
- प्रश्न-पत्र में दाहिने हाथ की ओर दिए गए प्रश्न-पत्र कोड को परीक्षार्थी उत्तर-पुस्तिका के मुख-पृष्ठ पर लिखें।
- कृपया जाँच कर लें कि इस प्रश्न-पत्र में 35 प्रश्न हैं।
- कृपया प्रश्न का उत्तर लिखना शुरू करने से पहले, उत्तर-पुस्तिका में प्रश्न का क्रमांक अवश्य लिखें।
- इस प्रश्न-पत्र को पढ़ने के लिए 15 मिनट का समय दिया गया है। प्रश्न-पत्र का वितरण पूर्वाह्न में 10.15 बजे किया जाएगा। 10.15 बजे से 10.30 बजे तक परीक्षार्थी केवल प्रश्न-पत्र को पढ़ेंगे और इस अवधि के दौरान वे उत्तर-पुस्तिका पर कोई उत्तर नहीं लिखेंगे।
- Please check that this question paper contains 23 printed pages.
- Q.P. Code given on the right hand side of the question paper should be written on the title page of the answer-book by the candidate.
- Please check that this question paper contains 35 questions.
- Please write down the serial number of the question in the answer-book before attempting it.
- 15 minute time has been allotted to read this question paper. The question paper will be distributed at 10.15 a.m. From 10.15 a.m. to 10.30 a.m., the candidates will read the question paper only and will not write any answer on the answer-book during this period.



56/2/2

245 B

1



P.T.O.

Read the following instructions very carefully and follow them :

- (i) This Question Paper contains **35** questions. **All** questions are compulsory.
- (ii) Question Paper is divided into **FIVE** sections – Section **A, B, C, D** and **E**.
- (iii) **In section A** – question number **1** to **18** are Multiple Choice (MCQ) type questions carrying **1** mark each.
- (iv) **In section B** – question number **19** to **25** are Very Short Answer (VSA) type questions carrying **2** marks each.
- (v) **In section C** – question number **26** to **30** are Short Answer (SA) type questions carrying **3** marks each.
- (vi) **In section D** – question number **31** & **32** are case-based questions carrying **4** marks each.
- (vii) **In section E** – question number **33** to **35** are Long Answer (LA) questions carrying **5** marks each.
- (viii) There is no overall choice. However, an internal choice has been provided in 2 questions in Section **B**, 2 questions in Section **C**, 2 questions in Section **D** and 2 questions in Section **E**.
- (ix) Use of calculator is **NOT** allowed.

1. Which of the following belongs to the class of Vinyl halides ? 1

- (a) $\text{CH}_2 = \text{CHCH}_2\text{CH}_2\text{Cl}$ (b) $\text{CH}_2 = \underset{\text{Br}}{\text{C}} - \text{CH}_3$
- (c) $\text{CH}_2 = \text{CH} - \text{CH}_2 - \text{Br}$ (d) $\text{CH} \equiv \text{C} - \text{Br}$

2. What is the secondary valency of Cobalt in $[\text{Co}(\text{en})_3\text{Cl}]^+$? 1

- (a) 6
- (c) 2
- (b) 4
- (d) 8

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3. When Benzene diazonium chloride reacts with phenol, it forms a dye. This reaction is called 1
- (a) Diazotisation reaction (b) Condensation reaction
(c) Coupling reaction (d) Acetylation reaction
4. The slope in the plot of $[R]$ vs. time for a zero order reaction is 1
- (a) $\frac{+k}{2.303}$ (b) $-k$
(c) $\frac{-k}{2.303}$ (d) $+k$
5. Proteins are polymers of 1
- (a) Nucleic acids (b) Amino acids
(c) Monosaccharides (d) Amines
6. Retention of configuration is observed in 1
- (a) S_N1 reaction
(b) S_N2 reaction
(c) Neither S_N1 nor S_N2 reaction
(d) S_N2 reaction as well as S_N1 reaction
7. An azeotropic mixture of two liquids will have a boiling point lower than either of the two liquids when it 1
- (a) shows a negative deviation from Raoult's law
(b) forms an ideal solution
(c) shows a positive deviation from Raoult's law
(d) is saturated



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8. Which of the following solutions of KCl will have the highest value of molar conductivity ? 1
- (a) 0.01 M (b) 1 M
(c) 0.5 M (d) 0.1 M
9. Which of the following does not give Cannizaro reaction ? 1
- (a) $(CH_3)_3C - CHO$ (b) $(CH_3)_2CH - CHO$
(c) -CHO (d) $HCHO$
10. Which of the following reactions are feasible ? 1
- (a) $CH_3CH_2Br + Na^+ O^-C(CH_3)_3 \rightarrow CH_3CH_2-O-C(CH_3)_3$
(b) $(CH_3)_3C - Cl + Na^+ O^-CH_2CH_3 \rightarrow CH_3CH_2-O-C(CH_3)_3$
(c) Both (a) and (b)
(d) Neither (a) nor (b)
11. Aldehydes and ketones react with hydroxylamine to form 1
- (a) hydrazones (b) cyanohydrins
(c) semicarbazones (d) Oxime
12. For a reaction $2A \rightarrow 3B$, rate of reaction $-\frac{d[A]}{dt}$ is equal to 1
- (a) $\frac{+3}{2} \frac{d[B]}{dt}$ (b) $\frac{+2}{3} \frac{d[B]}{dt}$
(c) $\frac{+1}{3} \frac{d[B]}{dt}$ (d) $+\frac{2d[B]}{dt}$



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13. The formula of the complex dichloridobis (ethane –1, 2-diamine) platinum (IV) nitrate is 1
- (a) $[\text{Pt Cl}_2(\text{en})_2 (\text{NO}_3)_2]$ (b) $[\text{Pt Cl}_2(\text{en})_2] (\text{NO}_3)_2$
(c) $[\text{Pt Cl}_2(\text{en})_2 (\text{NO}_3)]\text{NO}_3$ (d) $[\text{Pt} (\text{en})_2 (\text{NO}_3)_2]\text{Cl}_2$

14. Which one among the following metals of 3d series has the lowest melting point ? 1
- (a) Fe (b) Mn
(c) Zn (d) Cu

Given below are two statements labelled as Assertion (A) and Reason (R).
Select the most appropriate answer from the options given below :

- (a) Both (A) and (R) are true and (R) is the correct explanation of (A).
(b) Both (A) and (R) are true, but (R) is not the correct explanation of (A).
(c) (A) is true, but (R) is false.
(d) (A) is false, but (R) is true.
15. **Assertion (A) :** Elevation in boiling point is a colligative property. 1
Reason (R) : The lowering of vapour pressure of solution causes elevation in boiling point.
16. **Assertion (A) :** Chlorobenzene is resistant to electrophilic substitution reaction. 1
Reason (R) : C-Cl bond in chlorobenzene acquires partial double bond characters due to resonance.
17. **Assertion (A) :** Conductivity decreases with decrease in concentration of electrolyte. 1
Reason (R) : Number of ions per unit volume that carry the current in a solution decreases on dilution.
18. **Assertion (A) :** Transition metals have high enthalpy of atomisation. 1
Reason (R) : Greater number of unpaired electrons in transition metals results in weak metallic bonding.



SECTION – B

19. (a) Give reasons : **2 × 1**
- (i) Mercury cell delivers a constant potential during its life time.
 - (ii) In the experimental determination of electrolytic conductance, Direct Current (DC) is not used.

OR

- (b) Define fuel cell with an example. What advantages do the fuel cells have over primary and secondary batteries ? **2**
20. The vapour pressure of pure liquid X and pure liquid Y at 25 °C are 120 mm Hg and 160 mm Hg respectively. If equal moles of X and Y are mixed to form an ideal solution, calculate the vapour pressure of the solution. **2**

21. (a) Write the IUPAC names of the following : **2 × 1**
- (i) $[\text{Co}(\text{NH}_3)_5(\text{ONO})]^{2+}$
 - (ii) $\text{K}_2[\text{NiCl}_4]$

OR

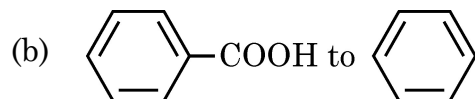
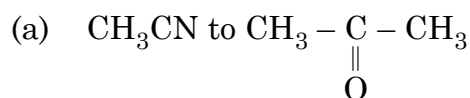
- (b) (i) What is a chelate complex ? Give one example.
- (ii) What are heteroleptic complexes ? Give one example. **2 × 1**
22. (a) The conversion of molecule A to B followed second order kinetics. If concentration of A increased to three times, how will it affect the rate of formation of B ? **2 × 1**
- (b) Define Pseudo first order reaction with an example.
23. (a) What is the difference between a nucleoside and nucleotide ? **2 × 1**
- (b) What products would be formed when a nucleotide from DNA containing thymine is hydrolysed ?

24. Write the chemical equation involved in the following : **2 × 1**
- (a) Kolbe's reaction
 - (b) Williamson synthesis



25. Do the following conversions in not more than two steps :

2 × 1



SECTION – C

26. (a) Draw the geometrical isomers of $[\text{Co}(\text{en})_2\text{Cl}_2]^{2+}$. Which geometrical isomer of $[\text{Co}(\text{en})_2\text{Cl}_2]^{2+}$ is not optically active and why ?

2 + 1

(b) Write the hybridisation and magnetic behaviour of $[\text{CoF}_6]^{3-}$.

[Given : Atomic number of Co = 27]

27. When 19.5 g of $\text{F} - \text{CH}_2 - \text{COOH}$ (Molar mass = 78 g mol⁻¹), is dissolved in 500 g of water, the depression in freezing point is observed to be 1°C. Calculate the degree of dissociation of $\text{F} - \text{CH}_2 - \text{COOH}$.

[Given : K_f for water = 1.86 K kg mol⁻¹]

3

28. Answer any 3 of the following :

3 × 1

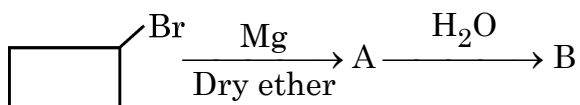
(a) Which isomer of C_5H_{10} gives a single monochloro compound $\text{C}_5\text{H}_9\text{Cl}$ in bright sunlight ?

(b) Arrange the following compounds in increasing order of reactivity towards $\text{S}_\text{N}2$ reaction :

2-Bromopentane, 1-Bromopentane, 2-Bromo-2-methylbutane



- (c) Why p-dichlorobenzene has higher melting point than those of ortho- and meta-isomers ?
- (d) Identify A and B in the following :



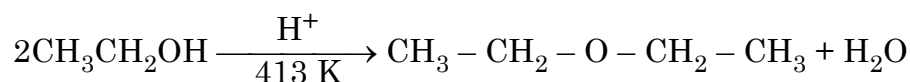
29. A first order reaction is 50% complete in 30 minutes at 300 K and in 10 minutes at 320 K. Calculate activation energy (E_a) for the reaction.

$$[R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}]$$

3

$$[\text{Given : } \log 2 = 0.3010, \log 3 = 0.4771, \log 4 = 0.6021]$$

30. (a) (i) Write the mechanism of the following reaction : 2 + 1



- (ii) Why ortho-nitrophenol is steam volatile while para-nitrophenol is not ?

OR

- (b) What happens when

- (i) Anisole is treated with CH_3Cl /anhydrous AlCl_3 ? 3 × 1
- (ii) Phenol is oxidised with $\text{Na}_2\text{Cr}_2\text{O}_7/\text{H}^+$?
- (iii) $(\text{CH}_3)_3\text{C} - \text{OH}$ is heated with $\text{Cu}/573 \text{ K}$?

Write chemical equation in support of your answer.



SECTION – D

The following questions are case based questions. Read the passage carefully and answer the questions that follow :

31. Carbohydrates are optically active polyhydroxy aldehydes and ketones. They are also called saccharides. All those carbohydrates which reduce Fehling's solution and Tollen's reagent are referred to as reducing sugars. Glucose, the most important source of energy for mammals, is obtained by the hydrolysis of starch. Vitamins are accessory food factors required in the diet. Proteins are the polymers of α -amino acids and perform various structural and dynamic functions in the organisms. Deficiency of vitamins leads to many diseases.

Answer the following :

- (a) The penta-acetate of glucose does not react with Hydroxylamine.
What does it indicate ? 1
- (b) Why cannot vitamin C be stored in our body ? 1
- (c) Define the following as related to proteins :
- (i) Peptide linkage
- (ii) Denaturation 2 × 1

OR

- (c) Define the following as related to carbohydrates :
- (i) Anomers
- (ii) Glycosidic linkage 2 × 1



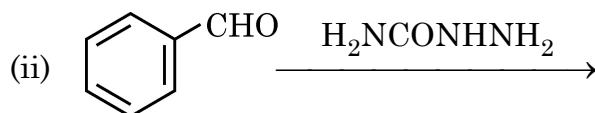
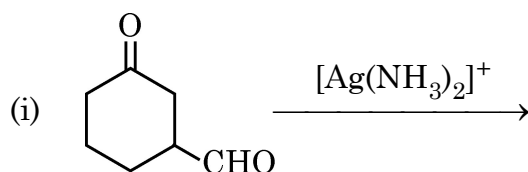
32. The carbon – oxygen double bond is polarised in aldehydes and ketones due to higher electronegativity of oxygen relative to carbon. Therefore they undergo nucleophilic addition reactions with a number of nucleophiles such as HCN, NaHSO₃, alcohols, ammonia derivatives and Grignard reagents. Aldehydes are easily oxidised by mild oxidising agents as compared to ketones. The carbonyl group of carboxylic acid does not give reactions of aldehydes and ketones. Carboxylic acids are considerably more acidic than alcohols and most of simple phenols.

Answer the following :

- (a) Write the name of the product when an aldehyde reacts with excess alcohol in presence of dry HCl. 1
- (b) Why carboxylic acid is a stronger acid than phenol ? 1
- (c) (i) Arrange the following compounds in increasing order of their reactivity towards CH₃MgBr :
- $$\text{CH}_3\text{CHO}, (\text{CH}_3)_3\text{C} - \underset{\text{O}}{\underset{\parallel}{\text{C}}} - \text{CH}_3, \text{CH}_3 - \underset{\text{O}}{\underset{\parallel}{\text{C}}} - \text{CH}_3$$
- (ii) Write a chemical test to distinguish between propanal and propanone. 2 × 1

OR

- (c) Write the main product in the following :



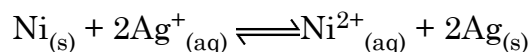
2 × 1



SECTION – E

33. (a) Conductivity of 2×10^{-3} M methanoic acid is 8×10^{-5} S cm⁻¹. Calculate its molar conductivity and degree of dissociation if $\wedge_m^0$ for methanoic acid is 404 S cm²mol⁻¹. **3 + 2**

- (b) Calculate the $\Delta_r G^\circ$ and $\log K_c$ for the given reaction at 298 K :



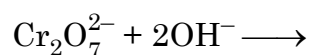
Given : $E^\circ_{\text{Ni}^{2+}/\text{Ni}} = -0.25$ V, $E^\circ_{\text{Ag}^+/\text{Ag}} = +0.80$ V

1F = 96500 C mol⁻¹.

34. (a) (I) Account for the following : **3 + 2**
- (i) E° value for $\text{Mn}^{3+}/\text{Mn}^{2+}$ couple is much more positive than that for $\text{Cr}^{3+}/\text{Cr}^{2+}$.
 - (ii) Sc^{3+} is colourless whereas Ti^{3+} is coloured in an aqueous solution.
 - (iii) Actinoids show wide range of oxidation states.
- (II) Write the chemical equations for the preparation of KMnO_4 from MnO_2 .

OR

- (b) (I) Account for the following : **2 + 2 + 1**
- (i) Transition metals form alloys.
 - (ii) Ce^{4+} is a strong oxidising agent.
- (II) Write one similarity and one difference between chemistry of Lanthanoids and Actinoids.
- (III) Complete the following ionic equation :



35. (a) (I) Give reasons :

3 + 2

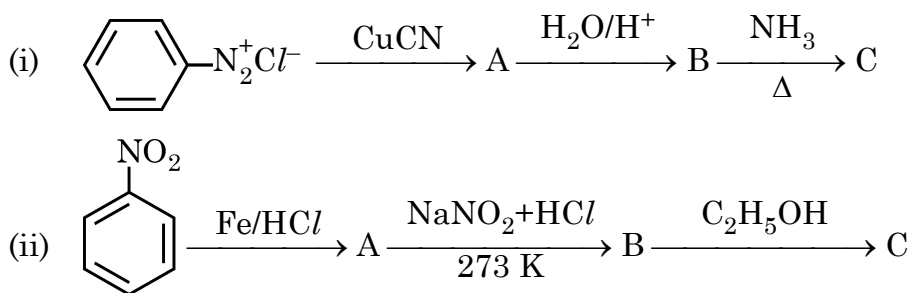
- (i) Aniline on nitration gives good amount of m-nitroaniline, though $-\text{NH}_2$ group is o/p directing in electrophilic substitution reactions.
- (ii) $(\text{CH}_3)_2\text{NH}$ is more basic than $(\text{CH}_3)_3\text{N}$ in an aqueous solution.
- (iii) Ammonolysis of alkyl halides is not a good method to prepare pure primary amines.

(II) Write the reaction involved in the following :

- (i) Carbyl amine test
- (ii) Gabriel phthalimide synthesis

OR

(b) (I) Write the structures of A, B and C in the following reactions : 3 + 1 + 1



(II) Why aniline does not undergo Friedal-Crafts reaction ?

(III) Arrange the following in increasing order of their boiling point :

