

Series DAA1B/3

Set-2

प्रश्न-पत्र कोड
Q.P. Code

31/3/2

रोल नं.

Roll No.

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परीक्षार्थी प्रश्न-पत्र कोड को उत्तर-पुस्तिका के मुख-पृष्ठ पर अवश्य लिखें।

Candidates must write the Q.P. Code on the title page of the answer-book.

विज्ञान SCIENCE

निर्धारित समय : 3 घण्टे

Time allowed : 3 hours

अधिकतम अंक : 80

Maximum Marks : 80

नोट	NOTE
(I) कृपया जाँच कर लें कि इस प्रश्न-पत्र में मुद्रित पृष्ठ 31 हैं।	(I) Please check that this question paper contains 31 printed pages.
(II) कृपया जाँच कर लें कि इस प्रश्न-पत्र में 39 प्रश्न हैं।	(II) Please check that this question paper contains 39 questions.
(III) प्रश्न-पत्र में दाहिने हाथ की ओर दिए गए प्रश्न-पत्र कोड को परीक्षार्थी उत्तर-पुस्तिका के मुख-पृष्ठ पर लिखें।	(III) Q.P. Code given on the right hand side of the question paper should be written on the title page of the answer-book by the candidate.
(IV) कृपया प्रश्न का उत्तर लिखना शुरू करने से पहले, उत्तर-पुस्तिका में प्रश्न का क्रमांक अवश्य लिखें।	(IV) Please write down the serial number of the question in the answer-book before attempting it.
(V) इस प्रश्न-पत्र को पढ़ने के लिए 15 मिनट का समय दिया गया है। प्रश्न-पत्र का वितरण पूर्वाह्न में 10.15 बजे किया जाएगा। 10.15 बजे से 10.30 बजे तक परीक्षार्थी केवल प्रश्न-पत्र को पढ़ेंगे और इस अवधि के दौरान वे उत्तर-पुस्तिका पर कोई उत्तर नहीं लिखेंगे।	(V) 15 minute time has been allotted to read this question paper. The question paper will be distributed at 10.15 a.m. From 10.15 a.m. to 10.30 a.m., the candidates will read the question paper only and will not write any answer on the answer-book during this period.

General Instructions :

Read the following instructions very carefully and strictly follow them :

- (i) *This question paper consists of **39** questions. **All** questions are compulsory.*
- (ii) *This Question paper is divided into **five** sections – **A, B, C, D** and **E**.*
- (iii) ***Section A** - Question Nos. **1** to **20** are Multiple Choice Questions. Each question carries **1** mark.*
- (iv) ***Section B** - Question Nos. **21** to **26** are Very Short Answer type questions. Each question carries **2** marks. Answer to these questions should be in the range of **30** to **50** words.*
- (v) ***Section C** - Question Nos. **27** to **33** are Short Answer (SA) type questions. Each question carries **3** marks. Answer to these questions should be in the range of **50** to **80** words.*
- (vi) ***Section D** - Question Nos. **34** to **36** are Long Answer type questions. Each question carries **5** marks. Answer to these questions should be in the range of **80** to **120** words.*
- (vii) ***Section E** - Question Nos. **37** to **39** are of 3 source-based/case-based units of assessment carrying **4** marks each with sub-parts.*
- (viii) *There is no overall choice. However, an internal choice has been provided in some sections. Only one of the alternatives has to be attempted in such questions.*

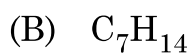
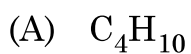
SECTION – A**(20 × 1 = 20)****In this section, Question Nos. 1 to 20 are Multiple-Choice Questions.****All questions are compulsory.**

1. A chemical compound used in glass, soap and paper industries is 1
(A) Washing Soda (B) Baking Soda
(C) Bleaching Powder (D) Common Salt
2. Which one of the following reactions is different from the remaining three ? 1
(A) $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{NaNO}_3$
(B) $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$
(C) $\text{KNO}_3 + \text{H}_2\text{SO}_4 \rightarrow \text{KHSO}_4 + \text{HNO}_3$
(D) $\text{ZnCl}_2 + \text{H}_2\text{S} \rightarrow \text{ZnS} + 2\text{HCl}$
3. Identify the product 'X' obtained in the following chemical reaction : 1
$$\text{CaCO}_3 \xrightarrow{\Delta} \text{'X'} + \text{CO}_2$$

(A) Quick lime (B) Gypsum
(C) Lime Stone (D) Plaster of Paris
4. Select a pair of natural indicator from the following : 1
(A) Litmus and methyl orange
(B) Turmeric and Litmus
(C) Phenolphthalein and methyl orange
(D) Methyl orange and Turmeric

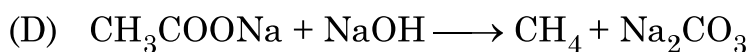
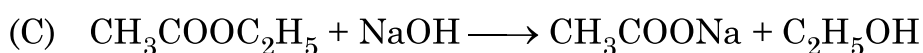
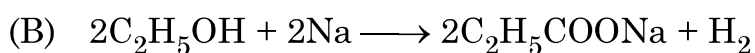
5. नीचे दिया गया कौन सा एक हाइड्रोकार्बन अन्य से भिन्न है ?

1



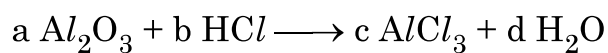
6. निम्नलिखित में से साबुनीकरण अभिक्रिया चुनिए :

1

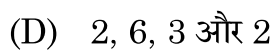
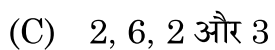
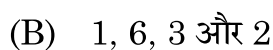
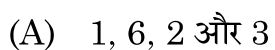


7. नीचे दिए गए रासायनिक समीकरण पर विचार कीजिए :

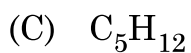
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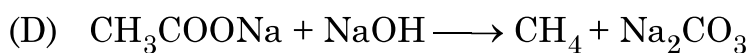
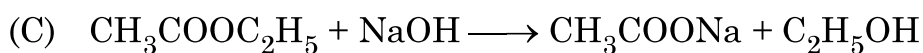
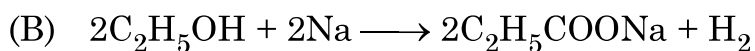
इस रासायनिक समीकरण को संतुलित बनाने के लिए a, b, c और d के मान क्रमशः होने चाहिए -



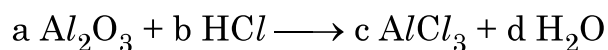
5. Which one of the following hydrocarbons is different from the others ? 1



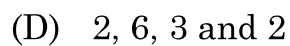
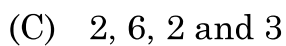
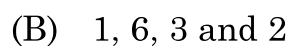
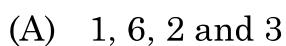
6. Select saponification reaction from the following : 1



7. Consider the following chemical equation : 1



In order to balance this Chemical equation, the values of a, b, c and d must be



8. In case of four wires of same material, the resistance will be minimum if the diameter and length of the wire respectively are 1
- (A) $D/2$ and $L/4$ (B) $D/4$ and $4L$
(C) $2D$ and L (D) $4D$ and $2L$
9. United Nations Environment Programme forged an agreement to 1
- (A) control CO_2 emissions in the environment
(B) conserve biodiversity
(C) control water pollution
(D) reduce CFC production
10. Which one of the following statements is TRUE for Hydra, Amoeba and Spirogyra ? 1
- (A) They are multicellular. (B) They are unicellular.
(C) They reproduce sexually. (D) They reproduce asexually.
11. A cross between two tall pea plants resulted in offsprings having a few dwarf plants. The gene-combination of the parental plants must be 1
- (A) Tt and Tt (B) Tt and tt
(C) TT and tt (D) TT and Tt

12. Which of the following statement (s) is (are) true about human heart ? 1

- (a) Right atrium receives oxygenated blood from lungs through pulmonary artery.
- (b) Left atrium transfers oxygenated blood to left ventricle which sends it to various parts of the body.
- (c) Right atrium receives deoxygenated blood from different parts of the body through vena cava.
- (d) Left atrium transfers oxygenated blood to aorta which sends it to different parts of the body.

- (A) (b) only (B) (a) and (d)
- (C) (b) and (c) (D) (b) and (d)

13. Part(s) of a flower which attracts insects for pollination is (are) 1

- (A) petals and Sepals (B) anther and Stigma
- (C) petals only (D) sepals only

14. The Phenomena of light involved in the formation of a rainbow in the sky are 1

- (A) Refraction, dispersion and reflection
- (B) Refraction, dispersion and total internal reflection
- (C) Dispersion, scattering and reflection
- (D) Dispersion, refraction and internal reflection

15. Consider the following statements about ozone :

1

- (a) Ozone is poisonous gas.
- (b) Ozone shields the earth's surface from the infrared radiation from the sun.
- (c) Ozone is a product of UV radiations acting on oxygen molecule.
- (d) At the lower level of the earth's atmosphere, ozone performs most essential function.

The correct statements are

- | | |
|-----------------|-----------------|
| (A) (a) and (b) | (B) (a) and (c) |
| (C) (b) and (c) | (D) (b) and (d) |

16. A food chain will be more advantageous in terms of energy if it has

1

- | | |
|----------------------|----------------------|
| (A) 2 trophic levels | (B) 3 trophic levels |
| (C) 4 trophic levels | (D) 5 trophic levels |

Q. 17 to 20 consists of two statements – Assertion (A) and Reason (R), answer these questions selecting the appropriate option given below :

- (A) Both (A) and (R) are true and (R) is the correct explanation of (A).
- (B) Both (A) and (R) are true and (R) is not correct explanation of (A).
- (C) (A) is true, but (R) is false.
- (D) (A) is false, but (R) is true.

17. **Assertion (A) :** Oxides of metals show basic characters.

1

Reason (R) : Oxides of metals react with acid to form salt and water.

18. **Assertion (A) :** Red light signals are used to stop the vehicles on the road. **1**

Reason (R) : Red coloured light is scattered the most so as to be visible from a large distance.

19. **Assertion (A) :** In human beings, males have 'XX' sex chromosomes and females have 'XY' sex chromosomes. **1**

Reason (R) : Sex of the child is determined at the time of fertilisation when male and female gamete fuse to form a zygote.

20. **Assertion (A) :** Receptors are usually located in our sense organs and perceive a particular stimulus. **1**

Reason (R) : Different sense organs have different receptors for detecting stimuli.

SECTION – B

Question Nos. 21 to 26 are very short answer type questions. Each question carries 2 marks.

21. When a few drops of Barium chloride solution are added to an aqueous solution of Sodium sulphate, a white precipitate is obtained. **2**

(a) Write balanced chemical equation for the reaction involved.

(b) What is the other name of this precipitation reaction ? Why is it called so ?

22. When do we say that a particular person is suffering from hypermetropia ? List two causes of this defect. Name the type of lens used to correct this defect. **2**

- 
23. (a) Draw a labelled diagram to show the pattern of magnetic field lines produced due to a current carrying straight conductor. Mark on it the direction of current in the conductor and the direction of magnetic field lines. 2

OR

23. (b) Name the device used to magnetise a piece of magnetic material. Draw a labelled diagram to show the arrangement used for the magnetisation of a cylinder made of soft iron. 2

24. Where are auxins synthesized ? How do they promote phototropism ? 2

25. (a) List any two pairs of visible contrasting characters of garden pea plants used by Mendel for his experiments stating the dominant and recessive characters in each pair. 2

OR

25. (b) In human beings, the probability of getting a male or a female child is 50%. Explain with the help of a flow diagram only. 2

26. Define the terms biodegradable and non-biodegradable substance. Classify the following items into these two categories : 2

Newspapers, Glass bottles, Polythene bags, Vegetable peels



SECTION – C

Question Nos. 27 to 33 are short answer type questions. Each question carries 3 marks.

27. (a) State giving reason the reduction process to obtain the following metals from their compounds : **3**

(i) Mercury,

(ii) Copper and

(iii) Sodium

OR

27. (b) State giving reason for the change in appearance observed when each of the following metal is exposed to atmospheric air for some time : **3**

(i) Silver,

(ii) Copper and

(iii) Iron

28. Write electronic configuration of Sodium (At. No.11) and Oxygen (At. No. 8) and show the formation of the ionic compound obtained when these two elements combine. Name anion and cation present in the compound. **3**

29. (a) List two constituents of Central Nervous System (CNS). How are these components protected from injuries ?

(b) Write two limitations of the use of electrical impulses. **3**

30. Write the main difference between aerobic and anaerobic respiration. State the pathway which is common for both. Write the overall chemical equation of aerobic respiration and mention the site where this process occurs inside the cells. 3
31. Explain in brief the function of an electric fuse in a domestic circuit. An electric heater of current rating 3 kW; 220 V is to be operated in an electric circuit of rating 5 A. What is likely to happen when the heater is switched 'ON' ? Justify your answer with necessary calculation. 3
32. Name and explain the phenomenon of light due to which the path of a beam of light becomes visible when it enters a smoke filled room through a small hole. Also state the dependence of colour of the light we receive on the size of the particle of the medium through which the beam of light passes. 3
33. List three advantages of parallel circuits in domestic wiring. 3

SECTION – D

Question Nos. 34 to 36 are long answer type questions. Each question carries 5 marks.

34. (a) (i) Draw a ray diagram to show the path of the refracted ray in each of the following cases :
- A ray of light incident on a concave lens
- (1) parallel to its principal axis, and
 - (2) is directed towards its principal focus.

- (ii) A 4 cm tall object is placed perpendicular to the principal axis of convex lens of focal length 24 cm. The distance of object from the lens is 16 cm. Find the position and size of image formed. **5**

OR

34. (b) (i) Draw a ray diagram to show the path of the reflected ray in each of the following cases :

A ray of light incident on a convex mirror

- (1) parallel to its principal axis, and
- (2) is directed towards its principal focus

- (ii) A 1.5 cm tall candle flame is placed perpendicular to the principal axis of a concave mirror of focal length 12 cm. If the distance of the flame from the pole of the mirror is 18 cm, use mirror formula to determine the position and size of the image formed. **5**

35. (a) What is puberty ? Write any two changes that occur in boys during early teenage years.

- (b) List two functions performed by testis in human males.

Mention one role each of (i) Vas deferens, (ii) Seminal Vesicle (iii) Urethra and (iv) Scrotum in human male reproductive system. **5**

OR

- (a) Write two functions each of the following parts in human female reproductive system :

- (i) Ovary
- (ii) Oviduct
- (iii) Uterus

- (b) Describe the structure and function of placenta. **5**

36. (a) (i) Five solutions A, B, C, D and E when tested with pH paper showed pH as 4, 1, 13, 7 and 10 respectively. Which solution is :

(1) Strongly acidic (2) Strongly alkaline (3) Weakly acidic
(4) Neutral and (5) Weakly alkaline ? Arrange the solutions in increasing order of H^+ ion concentration.

(ii) Write the name and formula of (1) an acidic salt and (2) a basic salt giving the name of the parent acid and parent base used to form the salt in each case.

5

OR

36. (b) Name and state in brief the process which is used to prepare sodium hydroxide from sodium chloride. In this process along with the main product two gases 'X' and 'Y' are also given off at the two electrodes. Name 'X' and 'Y' specifying the name of their respective electrode at which each gas is obtained. One of these gases when reacts with dry calcium hydroxide produces a compound 'Z' which is widely used in water treatment plants and textile industries. Name Z and write chemical equation for the reaction involved in its formation.

5

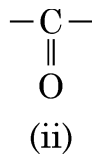
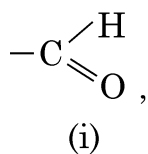
SECTION – E

Question Nos. 37 to 39 are Case/Source based questions. Each question carries 4 marks.

37. More than three million carbon compounds have been discovered in the field of chemistry. The diversity of these compounds is due to the capacity of carbon atoms for bonding with one another as well as with other atoms. Most of the carbon compounds are poor conductors of electricity and have low melting and boiling points. 4

(a) Write the molecular formula of first two members of homologous series having functional group –Br. 1

(b) Given below are the formulae of some functional groups : 1



Write the name of these functional groups.

- (c) What would be observed on adding a 5% alkaline potassium permanganate drop by drop to some warm ethanol taken in a test tube ? State the role of KMnO_4 in the reaction and write the chemical equation for the reaction involved. 2

OR

- (c) Write the name of the compound formed when ethanol is heated at 443 K temperature with excess of conc. H_2SO_4 . What is the role of conc. H_2SO_4 in the reaction ? Write the chemical equation for the reaction involved. 2

38. Human digestive system is a tube running from mouth to anus. Its main function is to breakdown complex molecules present in the food which cannot be absorbed as such into smaller molecules. These molecules are absorbed across the walls of the tube and the absorbed food reaches each and every cell of the body where it is utilised for obtaining energy. 4

(a) Name the glands present in the buccal cavity and write the components of food on which the secretion of these glands act upon. 1

(b) Two organs have a sphincter muscle at their exit. Name them. 1

(c) What will happen if :

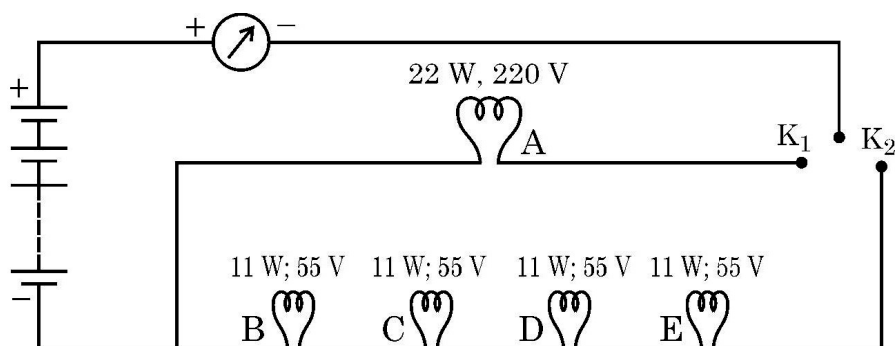
(i) mucus is not secreted by the gastric glands.

(ii) Villi are absent in the small intestine. 2

OR

(c) "Bile juice does not contain any enzyme, yet it has important roles in digestion." Justify the statement. 2

39. In a domestic circuit five LED bulbs are arranged as shown. The source voltage is 220 V and the power rating of each bulb is marked in the circuit diagram. Based on the following circuit diagram, answer the following questions : 4



(a) State what happens when 1

(i) key K_1 is closed.

(ii) key K_2 is closed.

(b) जब बल्ब B चमक रहा होता है तब वह कितनी धारा लेता है ?

1

(c) परिकल्पित कीजिए :

2

(i) बल्ब B का प्रतिरोध, तथा

(ii) चारों बल्ब B, C, D और E के संयोजन का कुल प्रतिरोध

अथवा

(c) परिपथ के सभी बल्बों की चमक का क्या होगा जबकि कुंजी K_1 तथा K_2 दोनों बन्द हैं तथा अचानक बल्ब C फ्यूज़ हो जाता है ? अपने उत्तर की पुष्टि के लिए कारण दीजिए ।

2



(b) Find the current drawn by the bulb B when it glows. **1**

(c) Calculate **2**

(i) the resistance of bulb B, and

(ii) total resistance of the combination of four bulbs B, C, D and E.

OR

(c) What would happen to the glow of all the bulbs in the circuit when keys K_1 and K_2 both are closed and the bulb C suddenly get fused ?

Give reason to justify your answer. **2**

